

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

- | | | | |
|------------|-----------|--------|------------------------|
| negative | noble gas | octet | electron configuration |
| positive | electrons | ions | polyatomic |
| crisscross | anion | cation | empirical formula |

1. An ionic bond is an attraction between oppositely charged _____.
2. Anions have a _____ charge.
3. An atom becomes an ion by losing or gaining _____.
4. The _____ rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) _____.
6. The atoms that make up _____ ions are bonded together by covalent bonds.
7. The _____ method can be used to write the formula for an ionic compound.
8. The _____ of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

9. Calcium becomes a monatomic cation by gaining two electrons.

10. A cation has a positive charge.

11. A binary ionic compound contains only one kind of cation and one kind of anion.

12. Monatomic ions consist of more than one atom.

13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons.

14. A great deal of energy is produced when an ionic compound is formed.

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