

Chapter 5 Reading Notes

Sunday, September 25, 2022

3:50 PM

- Octet rule: most main group elements will gain, lose, or share electrons to achieve a full valence shell of eight electrons
- In ionic compounds metallic elements typically lose electrons to fill a valence shell and non-metallic elements gain electrons to achieve a full valence shell
- In covalent compounds atoms seek to share electrons
 - Typically occurs between non-metallic elements
- Shared pair of electrons depicted as a straight line between two atoms in a Lewis Dot Diagram
- Bonds with unequal sharing of electrons are described as polarized because there is a negative end and positive pole of the molecule which results in an electric dipole
- Dipole moment: magnitude of the charge separation
 - Often represented with a special arrow (\rightarrow) where the + is oriented over the less electronegative atom and the arrowhead points to the more electronegative atom
 - Measured in units of debye
 - $1\text{debye}=3.34 \cdot 10^{-30}\text{C}\cdot\text{m}$
- Electron density often represented using colored contour maps called electrostatic potential diagrams
 - Regions with low electron density are colored blue and high ones are red
- Resonance structures are Lewis drawings that differ only in the position of bonding and non-bonding electrons not in the connectivity of the atoms
 - Structures in resonance are denoted by double headed arrows
- Isomer: molecules that have the same composition of atoms but different arrangement
- Octet rule can be used to predict which isomer will be most stable
- Atom with the lowest electronegativity is oftentimes the central atom and will never be hydrogen
- To find formal charge subtract the mu, number of non-bonding electrons and half the number of bonded electrons from the number of its valence electrons
- VSEPR method based on the fact that electrons being of like charge, repel each other and thus seek to be as far apart as possible

Chapter 5 Reading Notes Section 5 1 First Derivatives And

Katrin Zwirgmaier



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