

## CHAPTER 6

# Thermochemistry

### CHAPTER TERMS AND DEFINITIONS

Numbers in parentheses after definitions give the text sections in which the terms are explained. Starred terms are italicized in the text. Where a term does not fall directly under a text section heading, additional information is given for you to locate it.

**thermodynamics\*** science of the relationships between heat and other forms of energy (6.1, introductory section)

**thermochemistry\*** one area of thermodynamics; study of the quantity of heat absorbed or evolved by chemical reactions (6.1, introductory section)

**energy** potential or capacity to move matter (6.1)

**kinetic energy ( $E_k$ )** energy associated with an object by virtue of its motion (6.1)

**joule (J)** SI unit of energy,  $\text{kg} \cdot \text{m}^2/\text{s}^2$  (6.1)

**watt\*** measure of quantity of energy used per unit time;  $1 \text{ J/s}$  (6.1)

**calorie (cal)** non-SI unit of energy commonly used by chemists; originally defined as the amount of energy required to raise the temperature of one gram of water by one degree Celsius;  $1 \text{ cal} = 4.184 \text{ J}$  (exact definition) (6.1)

**potential energy ( $E_p$ )** an object's energy because of its position in a field of force (6.1)

**internal energy ( $U$ )** sum of kinetic and potential energies of the particles making up a substance (6.1)

**$E_{\text{tot}}$ \*** total energy of a substance; sum of the kinetic, potential, and internal energies of the substance (6.1)

**law of conservation of energy** energy may be converted from one form to another, but the total quantity of energy remains constant (6.1)

**thermodynamic system (or system)** substance or mixture of substances under study in which a physical or chemical change occurs (6.2)

**surroundings** everything in the vicinity of a thermodynamic system (6.2)

**heat ( $q$ )** energy that flows between system and surroundings because of a difference in temperature between the thermodynamic system and its surroundings (6.2)

**thermal equilibrium\*** state in which energy does not flow as heat between system and surroundings; temperature equality (6.2)

**heat of reaction** value of  $q$  required, at a given temperature, to return a system to the given temperature at the completion of the reaction (6.2)

**exothermic process** chemical reaction or physical change in which heat is evolved ( $q$  is negative) (6.2)

# Chapter 6 Thermochemistry Test

**SA Adler**



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