



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 5.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - CH_4
 - H_2SO_4
- Write the formula for the following compounds:
 - Carbon tetrachloride
 - Carbon disulfide
 - Carbon dioxide
 - Carbon monoxide

- Write the molecular formula for each of the following compounds:
 - Carbon tetrachloride
 - Carbon disulfide
 - Carbon dioxide
 - Carbon monoxide
- Write the molecular formula for each of the following compounds:
 - Carbon tetrachloride
 - Carbon disulfide
 - Carbon dioxide
 - Carbon monoxide
- Write the molecular formula for each of the following compounds:
 - Carbon tetrachloride
 - Carbon disulfide
 - Carbon dioxide
 - Carbon monoxide

SECTION 5.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw the Lewis structure for carbon tetrachloride. CCl_4
- Draw the Lewis structure for carbon dioxide. CO_2
- Draw the Lewis structure for ethane. $\text{H}-\text{C}-\text{C}-\text{H}$
- Draw the Lewis structure for acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$

- Draw the Lewis structure for each of the following compounds:
 - Carbon tetrachloride
 - Carbon disulfide
 - Carbon dioxide
 - Carbon monoxide

SECTION 5.3 BONDING: THE ORBITALS

- Which orbital and suborbital are represented by each of the following?
 - $2s$
 - $2p$
 - $3s$
 - $3p$
 - $4s$
 - $4p$
- Which orbital and suborbital are represented by each of the following?
 - $2s$
 - $2p$
 - $3s$
 - $3p$
 - $4s$
 - $4p$

Chapter 8 Covalent Bonding Test B Answers

Puja Mehta



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