

Name: _____ Date: _____ Period: _____

Periodic Trends Practice Worksheet

1. What is the difference between electronegativity and ionization energy?
2. Why does fluorine have a higher ionization energy than iodine?
3. Why do elements in the same family generally have similar properties?
4. Which is the largest atom in Group 4a (14)?
5. Which is the smallest atom in Group 7a (17)?
6. Which is the smallest atom in period 5?
7. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
8. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
9. Arrange the following atoms in order of decreasing atomic radius. Na Al P Cl Mg
10. For each of the following pairs, circle the element that is larger?
a) N^{3-} or F^{-1} b) Mg^{2+} or Ca^{2+} c) Fe^{2+} or Fe^{3+}
11. Circle the element in each pair has the larger radius?
a) Mg or Mg^{2+} b) O or O^{2-} c) K^{+} or Cl d) P^{3-} or S^{2-}
12. In each of the following pairs, circle the species with the higher first ionization energy:
a) Li or Cs (b) Cl or Ar (c) Ca or Br (d) Na or Ne (e) B or Be
13. In each of the following pairs, circle the species with the larger atomic radius:
a) Mg or Ba (b) S or S^{2-} (c) Cu^{+2} or Cu (d) He or H (e) Na or Cl
14. Circle the best choice in each list:
a) highest first ionization energy: C, N, Si
b) highest electronegativity: As, Sn, S
c) largest radius: S^{2-} , Cl^{-} , Cl
15. Order the following groups from largest to smallest radii.
a) Ar, Cl, K, S^{2-}
b) C, Al, F, Si
c) Na, Mg, Ar, P
d) I^{-} , Ba^{2+} , Cs^{+1} , F^{-1}

Chemistry Practice Test Periodic Trends And Orbitals

David Baud



Chemistry Practice Test Periodic Trends And Orbitals:

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Table of Contents Chemistry Practice Test Periodic Trends And Orbitals

1. Understanding the eBook Chemistry Practice Test Periodic Trends And Orbitals
 - The Rise of Digital Reading Chemistry Practice Test Periodic Trends And Orbitals
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practice Test Periodic Trends And Orbitals
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Practice Test Periodic Trends And Orbitals
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practice Test Periodic Trends And Orbitals
 - Personalized Recommendations
 - Chemistry Practice Test Periodic Trends And Orbitals User Reviews and Ratings
 - Chemistry Practice Test Periodic Trends And Orbitals and Bestseller Lists

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 - Chemistry Practice Test Periodic Trends And Orbitals eBook Subscription Services
 - Chemistry Practice Test Periodic Trends And Orbitals Budget-Friendly Options
6. Navigating Chemistry Practice Test Periodic Trends And Orbitals eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Practice Test Periodic Trends And Orbitals Compatibility with Devices
 - Chemistry Practice Test Periodic Trends And Orbitals Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Practice Test Periodic Trends And Orbitals
 - Highlighting and Note-Taking Chemistry Practice Test Periodic Trends And Orbitals
 - Interactive Elements Chemistry Practice Test Periodic Trends And Orbitals
8. Staying Engaged with Chemistry Practice Test Periodic Trends And Orbitals
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practice Test Periodic Trends And Orbitals
9. Balancing eBooks and Physical Books Chemistry Practice Test Periodic Trends And Orbitals
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practice Test Periodic Trends And Orbitals
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Practice Test Periodic Trends And Orbitals
 - Setting Reading Goals Chemistry Practice Test Periodic Trends And Orbitals
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Practice Test Periodic Trends And Orbitals
 - Fact-Checking eBook Content of Chemistry Practice Test Periodic Trends And Orbitals
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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