

Brønsted-Lowry Acids and Bases

1 Identify the hydrogen ion donor(s) and acceptor(s) in each of the following reactions:

	H ⁺ donors (the acids)	H ⁺ acceptors (the bases)
a) $\text{HNO}_3 (\text{aq.}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{H}_3\text{O}^+ (\text{aq.}) + \text{NO}_3^- (\text{aq.})$	_____	_____
b) $\text{C}_2\text{H}_5\text{NH}_2 (\text{aq.}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{C}_2\text{H}_5\text{NH}_3^+ (\text{aq.}) + \text{OH}^- (\text{aq.})$	_____	_____
c) $\text{CH}_3\text{COOH} (\text{aq.}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{CH}_3\text{COO}^- (\text{aq.}) + \text{H}_3\text{O}^+ (\text{aq.})$	_____	_____

2 Write the formulas for the conjugate base of each of the following acids.

a) H_2SO_3 _____	b) HCO_3^- _____	c) NH_4^+ _____
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3 Write the formulas for the conjugate acid of each of the following bases.

a) H_2O _____	b) CO_3^{2-} _____	c) PH_3 _____
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4 For each molecule or ion in the table, identify whether it can act as an acid or a base. Put a checkmark under each theory that describes it. There can be more than one theory per compound.

Molecule/ion	Acid or Base	Arrhenius	Brønsted-Lowry	Lewis
Br^-				
CN^-				
H_2CO_3				
NH_3				
HNO_2				
$\text{Ba}(\text{OH})_2$				
HCl				
AlCl_3				
Cl^-				
KOH				
IO_3^-				
CH_3COOH				

Chapter 14 Acids And Bases Worksheet

Clifford Lane



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Table of Contents Chapter 14 Acids And Bases Worksheet

1. Understanding the eBook Chapter 14 Acids And Bases Worksheet
 - The Rise of Digital Reading Chapter 14 Acids And Bases Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 14 Acids And Bases Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 14 Acids And Bases Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 14 Acids And Bases Worksheet

- Personalized Recommendations
 - Chapter 14 Acids And Bases Worksheet User Reviews and Ratings
 - Chapter 14 Acids And Bases Worksheet and Bestseller Lists
5. Accessing Chapter 14 Acids And Bases Worksheet Free and Paid eBooks
 - Chapter 14 Acids And Bases Worksheet Public Domain eBooks
 - Chapter 14 Acids And Bases Worksheet eBook Subscription Services
 - Chapter 14 Acids And Bases Worksheet Budget-Friendly Options
 6. Navigating Chapter 14 Acids And Bases Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 14 Acids And Bases Worksheet Compatibility with Devices
 - Chapter 14 Acids And Bases Worksheet Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 14 Acids And Bases Worksheet
 - Highlighting and Note-Taking Chapter 14 Acids And Bases Worksheet
 - Interactive Elements Chapter 14 Acids And Bases Worksheet
 8. Staying Engaged with Chapter 14 Acids And Bases Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 14 Acids And Bases Worksheet
 9. Balancing eBooks and Physical Books Chapter 14 Acids And Bases Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 14 Acids And Bases Worksheet
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 14 Acids And Bases Worksheet
 - Setting Reading Goals Chapter 14 Acids And Bases Worksheet
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 14 Acids And Bases Worksheet

- Fact-Checking eBook Content of Chapter 14 Acids And Bases Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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