

CHAPTER 14 REVIEW*Acids and Bases***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Name the following compounds as acids:

_____ sulfuric acid _____	a. H_2SO_4
_____ sulfurous acid _____	b. H_2SO_3
_____ hydrosulfuric acid _____	c. H_2S
_____ perchloric acid _____	d. HClO_4
_____ hydrocyanic acid _____	e. hydrogen cyanide

2. _____
- H_2S
- _____ Which (if any) of the acids mentioned in item 1 are binary acids?

3. Write formulas for the following acids:

_____ HNO_2 _____	a. nitrous acid
_____ HBr _____	b. hydrobromic acid
_____ H_3PO_4 _____	c. phosphoric acid
_____ CH_3COOH _____	d. acetic acid
_____ HClO _____	e. hypochlorous acid

4. Calcium selenate has the formula
- CaSeO_4
- .

_____ H_2SeO_4 _____	a. What is the formula for selenic acid?
_____ H_2SeO_3 _____	b. What is the formula for selenous acid?

5. Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.

6. Write balanced chemical equations for the following reactions of acids and bases:

- a. aluminum metal with dilute nitric acid



- b. calcium hydroxide solution with acetic acid



Chapter 14 Review Acids Bases Section 1 Modern Chemistry

Michael Brown



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