

Calculating Specific Heat

An iron skillet has a mass of 500.0 grams. The specific heat of iron is $0.449 \text{ J/g}^\circ\text{C}$. How much heat must be absorbed to raise the skillet's temperature by 95.0°C ?

1 Read and Understand

What information are you given?

Mass of iron, $m = 500.0 \text{ g}$

Specific heat of iron, $c = 0.449 \text{ J/g}^\circ\text{C}$

Temperature change, $\Delta T = 95.0^\circ\text{C}$

2 Plan and Solve

What unknown are you trying to calculate?

Amount of heat needed, $Q = ?$

What formula contains the given quantities and the unknown?

$$Q = m \times c \times \Delta T$$



Replace each variable with its known value.

$$\begin{aligned} Q &= 500.0 \text{ g} \times 0.449 \text{ J/g}^\circ\text{C} \times 95.0^\circ\text{C} \\ &= 21,375 \text{ J} = 21.4 \text{ kJ} \end{aligned}$$

3 Look Back and Check

Is your answer reasonable?

Round off the data to give a quick estimate.

$$Q = 500 \text{ g} \times 0.5 \text{ J/g}^\circ\text{C} \times 100^\circ\text{C} = 25 \text{ kJ}$$

This is close to 21.4 kJ, so the answer is reasonable.

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Victor M. Corman



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