

1. How many moles of water are produced with excess oxygen gas reacts with 3.21 moles of CH_4 ?



$$\frac{3.21 \text{ mol CH}_4}{1} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol CH}_4} = 6.42 \text{ mol H}_2\text{O}$$

2. How many moles of copper are needed to create 4.51 moles of Cu_2O when reacted with excess oxygen?



$$\frac{4.51 \text{ mol Cu}_2\text{O}}{1} \times \frac{2 \text{ mol Cu}}{1 \text{ mol Cu}_2\text{O}} = 9.02 \text{ mol Cu}$$

3. How many moles of water will be made when 5.33 moles of NaOH react completely with CO_2 ?



$$\frac{5.33 \text{ mol NaOH}}{1} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol NaOH}} = 2.67 \text{ mol H}_2\text{O}$$

4. If 32.1 moles of nitrogen gas react completely with H_2 , how many moles of NH_3 will be produced?



$$\frac{32.1 \text{ mol N}_2}{1} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol N}_2} = 64.2 \text{ mol H}_2\text{O}$$

5. How many moles of water will be produced if 21.1 moles of H_2O_2 decompose?



$$\frac{21.1 \text{ mol H}_2\text{O}_2}{1} \times \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}_2} = 21.1 \text{ mol H}_2\text{O}$$

6. How many moles of NaCl will be produced if 1.23 moles of BaCl_2 react completely with Na_2SO_4 ?



$$\frac{1.23 \text{ mol BaCl}_2}{1} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol BaCl}_2} = 2.46 \text{ mol H}_2\text{O}$$

7. How many moles of Si_2H_6 are needed to react with excess O_2 to produce 32.1 moles of SiO_2 ?



$$\frac{32.1 \text{ mol SiO}_2}{1} \times \frac{4 \text{ mol Si}_2\text{H}_6}{8 \text{ mol SiO}_2} = 16.05 \text{ mol Si}_2\text{H}_6$$

Chapter 3 The Mole And Stoichiometry Part 1 The Mole

Michael Hoelscher



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