

6. The new penny is lighter because it is mostly zinc which has a lower density than copper.

### You're The Chemist

- at 90°, mass of 1 drop: 0.019 g  
at 45°, mass of 1 drop: 0.0218 g  
at 0°, mass of 1 drop: 0.0242 g  
Pipets give different results.
- The best angle is 90° because the pipet is easiest to control. Expel the air bubble so that the first drop will be the same size as the rest.
- Find mass of can and divide by density of aluminum. Sample answer: mass of one can: 14.77 g;  
 $d_{Al}$ : 2.70 g/cm<sup>3</sup>;  $V = 5.47 \text{ cm}^3$
- (1) Measure the mass before and after you fill the can with water. Use the mass and density of water to find the volume.  
(2) Measure the height and radius and calculate volume.  
 $V = \pi r^2 h$   
(Can is not a perfect cylinder.)  
(3) Read label: 12 oz = 355 mL.
- Sample answer:  
 $V = 16.5 \text{ m} \times 3.0 \text{ m} \times 12.8 \text{ m}$   
 $= 630 \text{ m}^3 \times 1000 \text{ L/m}^3$   
 $= 630,000 \text{ L}$   
Assume 30 people with an average weight of 130 lb (1 kg = 2.2 lb) and a density of about 1.0 kg/L.  
Volume of 30 people:  
 $V = 30 \times 130 \text{ lb} \times 1 \text{ kg}/2.2 \text{ lb} \times 1 \text{ L}/1.0 \text{ kg}$   
 $= 1800 \text{ L}$   
The volume of 30 chairs, 15 tables, and 2 desks is about that of 30 people or 1800 L.  
The volume of people and furniture is 3600 L.  
 $\% \text{ error} = (3600 \text{ L}/630,000 \text{ L})(100\%)$   
 $= 0.57\%$ .
- If die measures 1.55 cm on a side:  
 $V = (1.55 \text{ cm})^3 = 3.72 \text{ cm}^3$   
A die has 21 holes that are hemispheres with a radius of 0.20 cm.  
Volume of hemisphere:  $2/3\pi r^3$   
 $V = 2/3\pi r^3 = 0.017 \text{ cm}^3$   
Volume of 21 hemispheres:  $0.36 \text{ cm}^3$   
Volume of die:

$$V = 3.72 \text{ cm}^3 - 0.36 \text{ cm}^3$$

$$= 3.36 \text{ cm}^3$$

$$\text{Error} = 0.36 \text{ cm}^3$$

Percent error

$$= (0.36 \text{ cm}^3/3.36 \text{ cm}^3)(100\%) = 11\%$$

Note: The holes in some dice are cones. The volume of a cone is  $1/3 \pi r^2 h$ .

- Find weight in pounds and convert to kg. Assume the density is about 1.00 kg/L.  
 $V = \text{weight} \times 1 \text{ kg}/2.2 \text{ lb} \times 1 \text{ L}/1.00 \text{ kg}$

## Section Review 4.1

### Part A Completion

- |              |                                     |
|--------------|-------------------------------------|
| 1. atoms     | 3. compounds                        |
| 2. different | 4. separated, joined, or rearranged |

### Part B True-False

- |       |       |       |
|-------|-------|-------|
| 5. NT | 6. ST | 7. AT |
|-------|-------|-------|

### Part C Matching

- |      |       |
|------|-------|
| 8. c | 10. d |
| 9. a | 11. b |

### Part D Questions and Problems

- simple whole-number ratios
- $\frac{1 \times 10^{23} \text{ atoms}}{\text{cm}^3} \times \frac{100 \text{ cm}^3}{\text{m}} = 10^{25} \text{ atoms/m}$

## Section Review 4.2

### Part A Completion

- |              |                      |
|--------------|----------------------|
| 1. subatomic | 5. nucleus           |
| 2. negative  | 6. Ernest Rutherford |
| 3. protons   | 7. positive          |
| 4. neutrons  | 8. electrons         |

### Part B True-False

- |        |        |
|--------|--------|
| 9. NT  | 11. NT |
| 10. AT | 12. AT |

### Part C Matching

- |       |       |       |
|-------|-------|-------|
| 13. c | 15. e | 17. b |
| 14. a | 16. d |       |

### Part D Questions and Problems

- protons, neutrons

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**Nivaldo J. Tro**



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*Adhesive Bonding* L.H. Lee, 2013-06-29 For several years I have been responsible for organizing and teaching in the fall a short course on Fundamentals of Adhesion Theory Practice and Applications at the State University of New York at New Paltz Every spring I would try to assemble the most pertinent subjects and line up several capable lecturers for the course However there has always been one thing missing an authoritative book that covers most aspects of adhesion and adhesive bonding Such a book would be used by the participants

as a main reference throughout the course and kept as a sourcebook after the course had been completed On the other hand this book could not be one of those All you want to know about volumes simply because adhesion is an interdisciplinary and ever growing field For the same reason it would be very difficult for a single individual especially me to undertake the task of writing such a book Thus I relied on the principle that one leaves the truly monumental jobs to experts and I finally succeeded in asking several leading scientists in the field of adhesion to write separate chapters for this collection Some chapters emphasize theoretical concepts and others experimental techniques In the humble beginning we planned to include only twelve chapters However we soon realized that such a plan would leave too much ground uncovered and we resolved to increase the coverage After the book had evolved into thirty chapters we started to feel that perhaps our mission had been accomplished

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