

Key Ideas

- ▶ What is the difference between protons, neutrons, and electrons?
- ▶ What do atoms of an element have in common with other atoms of the same element?
- ▶ Why do isotopes of the same element have different atomic masses?
- ▶ What unit is used to express atomic mass?

Key Terms

proton
neutron
atomic number
mass number
isotope
unified atomic mass unit
mole

Why It Matters

Radioisotopes emit energy when they decay. To diagnose and treat diseases, doctors use this property of radioisotopes to track where in the body certain atoms go.

Less than 100 years after Dalton published his atomic theory, scientists determined that atoms consisted of smaller particles, such as the electron. In this section, you will learn more about the particles inside the atom.

What Is in an Atom?

Atoms are made up of various subatomic particles. To understand the chemistry of most substances, however, we need to study only three of these particles. ▶ **The three main subatomic particles are distinguished by mass, charge, and location in the atom.** Figure 1 compares these particles.

At the center of each atom is a small, dense nucleus. The nucleus is made of **protons**, which have a positive charge, and **neutrons**, which have no charge. Protons and neutrons are almost identical in size and mass. Moving around outside the nucleus is a cloud of very tiny, negatively charged electrons. The mass of an electron is much smaller than that of a proton or neutron.

Figure 1 Subatomic Particles

Particle	Charge	Mass (kg)	Location in the atom
Proton	+1	1.67×10^{-27}	in the nucleus
Neutron	0	1.67×10^{-27}	in the nucleus
Electron	-1	9.11×10^{-31}	outside the nucleus

proton (PROE) is a subatomic particle that has a positive charge and that is located in the nucleus of an atom.

neutron (NEE) is a subatomic particle that has no charge and that is located in the nucleus of an atom.

Focus

In this section, students compare the three basic subatomic particles (protons, electrons, neutrons). They learn how atomic number, mass number, and molar mass are used to represent properties of atoms.

Bellringer

Use the Bellringer transparency to prepare students for this section.

Demonstrate

Counting Large Numbers by Mass

Put students into groups. Each group will need a large container, identical items to fill the container (such as beans), and a balance.

Ask students to find the following: the mass of the empty container (m_{cont}), the mass of a few (x) of the items ($m_{x \text{ items}}$), and the mass of the container with all of the items in it (m_{tot}). Tell each group to use their data to calculate the number of small items without counting them. (First, find the mass of one

item by dividing the mass of a given number of items by that number.

Then find the mass of all of the items by subtracting the mass of the container from the total mass. Finally, the number of unknown items (n) equals the mass of all of the items divided by the mass of one item.) Tell students that in this section they will see how this method can be used to “count atoms.” **LE Kinesthetic**

Key Resources

 **Teaching Transparencies**
C8 Nucleus
TM12 Subatomic Particles
C9 Isotopes
TM13 Mole-Mass Conversion

 **Visual Concepts**
Atomic Number
Mass Number
Isotopes and Nuclides
Average Atomic Mass
The Mole

 **Datasheet**
Modeling Isotopes

 **Science Skills Worksheets**
Scientific Notation
Significant Figures
Dimensional Analysis

 **Math Skills Worksheets**
Conversion Factors
Converting Amount to Mass
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