

## Chapter Test A

### Chapter: Chemical Equations and Reactions

Do the work provided below the letter of the item or phrase that best completes each statement or best answers each question.

1. The mass number of an atom is the sum of the number of \_\_\_\_\_ and the number of \_\_\_\_\_ in the nucleus of the atom.
  - a. The number of protons; neutrons
  - b. Electrons; protons; neutrons
  - c. Electrons; protons; neutrons; positrons
  - d. None of the above; all three terms being defined incorrectly
2. An element described by chemical equation  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  is \_\_\_\_\_ .
  - a. carbon hydrocarbon or gas
  - b. carbon gas product
  - c. carbon or a reactant
  - d. none of the above
3. Which of the following relationships is always true for a balanced chemical equation?
  - a. The products of a reaction are equal in mass to the reactants (law of conservation of mass).
  - b. All of the coefficients are the same.
  - c. The number of molecules of each reactant is equal to the number of molecules of each product.
  - d. The number of molecules of each reactant is equal to the number of molecules of each product.
4. There is a negative enthalpy change for a chemical reaction if the \_\_\_\_\_ is \_\_\_\_\_ .
  - a. reactants; less than the products
  - b. product; less than the reactants
  - c. products; more than the reactants
  - d. all of the above
5. The number of chemical equations that can be balanced by inspection is \_\_\_\_\_ .
  - a. equal the number
  - b. equal the number of products
  - c. equal the number of reactants
  - d. equal the coefficients

# Chapter 9 Modern Chemistry Test

**Disha Experts**



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[The Good Student](#) Charles Jay Rop, 1996

**Laboratory Experiments to Accompany "Modern Chemistry,"** Fredus Nelson Peters, 1902

**Evaluation in Modern Education** Jacob Wayne Wrightstone, 1956

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[Elementary Modern Chemistry](#) Wilhelm Ostwald, Harry Wheeler Morse, 1909

**Modern Road Emulsions** Road Emulsion Association. Technical Committee, 1958

*Scientific Protocols for Fire Investigation* John J. Lentini, 2006-01-26. *Scientific Protocols for Fire Investigation* provides comprehensive coverage from historical, developmental, current, and practical perspectives. The author, uniquely qualified with years of experience in both on-site investigations and lab analyses, provides a resource that is unparalleled in depth and focus. The book is distinctive in that it not

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**Report of the New England Association of Chemistry Teachers** ... New England Association of Chemistry Teachers, 1931

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