

CH-101 2008 Assignment 10 Class

Name: _____

Number: _____

10. Atmospheric pressure at elevations of 1000 ft average about 0.70 atm. Consider this pressure to maintain air masses, lakes of mercury, and gas-filled and vacuum sealed flasks at equilibrium at 1000 feet mean sea level (Mount St. Helens)?

What is the physical explanation regarding water pressure in commercial scuba tanks provided constructed to a pressure greater than the required hydrostatic pressure at construction of 1000 ft?

11. Suppose a 1000 mL flask contains a volume of 2.0 g. Water surface, during the pressure at 1.0 atm, has a depth of 200 cm, which the pressure is 2.0 atm. What would happen to the volume of the flask? What if the end of the flask was in a liquid that went to the surface and was exposed to another flask? What would happen to the air the flask contained?

12. A volume of gas is confined to a cylinder with a freely moving piston, and heat is gradually added to the system. In which of the following situations would the amount of gas, volume of the cylinder, temperature be most likely to change?

- (a) The pressure will increase as heat is added; the volume will double.
- (b) The pressure will increase as heat is added; the volume will double.
- (c) The pressure will increase as heat is added; the volume will increase but not double.
- (d) The pressure will increase but something else will double.

13. Use Charles's law to compare the volume pressure and amount of gas for the conditions:

	(a)	(b)	(c)	(d)
Pressure	1.0 atm	0.5 atm	1.0 atm	0.5 atm
Volume	1.0 L	2.0 L	1.0 L	2.0 L
Temperature	273 K	546 K	273 K	546 K
Amount	1.0 mol	1.0 mol	2.0 mol	2.0 mol

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**P. W. J. M. Boumans, C. J. W. Brooks, C.
Burgess**

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