

## 10

## CHEMICAL QUANTITIES

## Chapter Test A

## A. Matching

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

## Column A

- \_\_\_\_\_ 1. standard temperature and pressure
- \_\_\_\_\_ 2. molar volume
- \_\_\_\_\_ 3. representative particle
- \_\_\_\_\_ 4. molar mass
- \_\_\_\_\_ 5. atomic mass
- \_\_\_\_\_ 6. Avogadro's number
- \_\_\_\_\_ 7. mole (mol)
- \_\_\_\_\_ 8. empirical formula
- \_\_\_\_\_ 9. percent composition
- \_\_\_\_\_ 10. molar mass

## Column B

- a. the volume occupied by a mole of any gas at STP (22.4 L)
- b. the mass in amu of an atom of an element
- c. the number of representative particles in a mole of a substance
- d. the percent by mass of each element in a compound
- e. the mass of a mole of a substance
- f. the formula that gives the lowest whole-number ratio of the elements in a compound
- g. an atom, a formula unit, or a molecule
- h. the mass in grams of a mole of any pure substance
- i.  $6.02 \times 10^{23}$  representative particles of a substance
- j.  $0^\circ\text{C}$  and 101.3 kPa

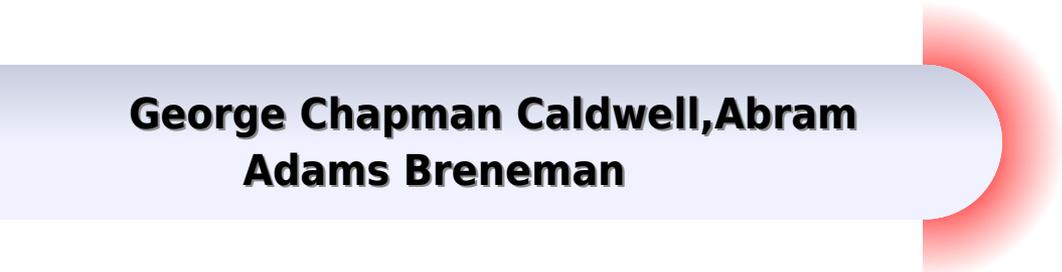
## B. Multiple Choice

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol,  $\text{C}_3\text{H}_7\text{O}$ ?
- a.  $4 \times (6.02 \times 10^{23})$
- b. 4
- c. 32
- d.  $32 \times (6.02 \times 10^{23})$
- \_\_\_\_\_ 12. The mass of a mole of NaCl is its:
- a. molar mass.
- b. empirical formula.
- c. gram atomic mass.
- d. atomic mass.

# Chemical Quantities Chapter 10 Test Answer Key

**George Chapman Caldwell, Abram  
Adams Breneman**



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