

# Mayya's Equilibrium Study Guide



$K_c$  is the Equilibrium Constant

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$



$K_p$  is the Equilibrium Constant  
when pressures are given

$$K_p = \frac{P_c^c P_d^d}{P_a^a P_b^b}$$

$$K_p = K_c(RT)^{\Delta n}$$

Where  $R=0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$ ,  $T$ = temperature in K,  
 $\Delta n$  = moles of gas products - moles of gas reactants

## How To Predict The Direction Of Reaction

$$Q = \frac{[C]^c \text{noneq} [D]^d \text{noneq}}{[A]^a \text{noneq} [B]^b \text{noneq}}$$

If  $Q > K$ , the reaction will go to the left  
If  $Q < K$ , the reaction will go to the right  
If  $Q = K$ , the reaction is at equilibrium

## ICE CHART

I = Initial (initial concentrations given in the problem)  
C = Change (in terms of X, including coefficients)  
E = Equilibrium (I+C)

Make an ICE Chart when you are asked to find equilibrium concentrations or pressures

## Equation Stoichiometry & Equilibrium Constant

- Reverse an equation = Take reciprocal of  $K$  ( $1/K$ )
- Multiply an equation by a constant = Raise  $K$  to the power of that constant
- Add equations = Multiply their  $K$ 's

# Chemical Equilibrium Study Guide Answers

**Wolfgang Guggemos**



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