

## Chemistry Worksheets Class 11 on Chapter 8 Redox Reactions - Set 1

**Q-1:** Which of the following statements about redox reactions is correct?

- a) A nonmetal with a higher reduction potential is a more potent reducing agent.
- b) A metal with a lower oxidation potential is a more potent oxidising agent.
- c) The oxygen oxidation state in  $O_2$  is -1.
- d) All of the above

**Q-2:** Which of the following can behave as a strong reducing agent?

- a)  $HClO_4$
- b)  $SnCl_2$
- c)  $SO_2$
- d) All of the above

**Q-3:** The pair of compounds in which both metals are in the highest oxidation state is

- a)  $FeSO_4$ ,  $CuS_2$
- b)  $CrO_2Cl_2$ ,  $MnO_4^-$
- c)  $TiO_2$ ,  $MnO_2$
- d)  $[Co(CN)_6]^{2-}$ ,  $MnO_2$

**Q-4:** The number of electrons involved when 1 mole of  $H_2O_2$  decomposes as

- $H_2O_2 \rightarrow H_2O + O_2$
- a) 1
  - b) 6
  - c) 8
  - d) 4

**Q-5:** Zinc iodide is formed during the reaction of zinc and iodine. Which of these is being oxidised?

- a) Zinc ions
- b) Zinc atom
- c) Iodide ions
- d) Iodine

**Q-6:** Can a solution of 1M copper sulphate be stored in a vessel made of nickel metal?

Given that  $E^\circ(Ni^{2+}/Ni) = -0.25$  volt and  $E^\circ(Cu^{2+}/Cu) = +0.34$  volt.

**Q-7:** Using the electronic concept, identify the oxidant and reductant in the following redox reactions.

- i)  $Zn + 2H^+ \rightarrow Zn^{2+} + H_2$
- ii)  $BrO_3^- + F_2 + 2OH^- \rightarrow BrO_4^- + 2F^- + H_2O$

# Chemistry Chapter 11 Chemical Reactions Packet Answers

**John McMurry, Susan McMurry**



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## **Table of Contents Chemistry Chapter 11 Chemical Reactions Packet Answers**

1. Understanding the eBook Chemistry Chapter 11 Chemical Reactions Packet Answers
  - The Rise of Digital Reading Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Chemistry Chapter 11 Chemical Reactions Packet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Personalized Recommendations
  - Chemistry Chapter 11 Chemical Reactions Packet Answers User Reviews and Ratings
  - Chemistry Chapter 11 Chemical Reactions Packet Answers and Bestseller Lists
5. Accessing Chemistry Chapter 11 Chemical Reactions Packet Answers Free and Paid eBooks
  - Chemistry Chapter 11 Chemical Reactions Packet Answers Public Domain eBooks
  - Chemistry Chapter 11 Chemical Reactions Packet Answers eBook Subscription Services
  - Chemistry Chapter 11 Chemical Reactions Packet Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 11 Chemical Reactions Packet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 11 Chemical Reactions Packet Answers Compatibility with Devices
  - Chemistry Chapter 11 Chemical Reactions Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Highlighting and Note-Taking Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Interactive Elements Chemistry Chapter 11 Chemical Reactions Packet Answers

8. Staying Engaged with Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 11 Chemical Reactions Packet Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 11 Chemical Reactions Packet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Setting Reading Goals Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Fact-Checking eBook Content of Chemistry Chapter 11 Chemical Reactions Packet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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