

KEY

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TEAM
1

Chemistry Lab Experiment

Determination of the Molar Volume of a Gas

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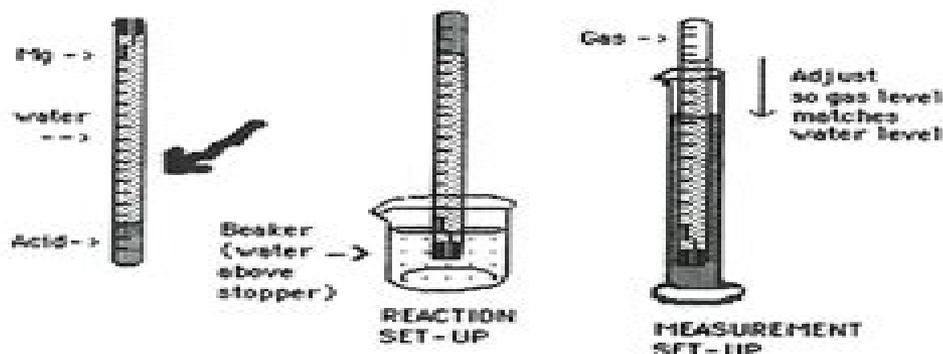
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Purpose:

- To experimentally determine the molar volume of hydrogen gas and to compare it with the $\frac{22.4\text{-L}}{\text{mole}}$ theoretical STP value.
- To relate gas law concepts learned in MA Frameworks Topic 6 (e.g., the Combined Gas Law and Dalton's Law of Partial Pressures) to previously learned concepts involving balancing chemical equations and stoichiometric relationships.

Introduction:



The **water displacement method** can be used to capture the hydrogen gas liberated when a strip of magnesium metal reacts with a $3.0\text{-}M$ solution of hydrochloric acid. The total gas volume of H_2O vapor and H_2 adjusted to the room's barometric pressure can be determined from the **eudiometer** (graduated gas-measuring tube) by partially immersing the tube into a large container of water until the liquid level inside the tube matches the water level in the large container.

According to **Dalton's Law of Partial Pressures**, the total gas pressure in the eudiometer is the sum of the pressure of the hydrogen gas collected plus the **water vapor pressure**. Symbolically, $P_{\text{total}} = P_{\text{H}_2} + P_{\text{H}_2\text{O}}$. Subtracting the water vapor pressure from the total pressure gives the pressure of the dry hydrogen gas.

Chemistry Molar Volume Of Hydrogen Lab Answers

L. Neal Carmichael, David F. Haines



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