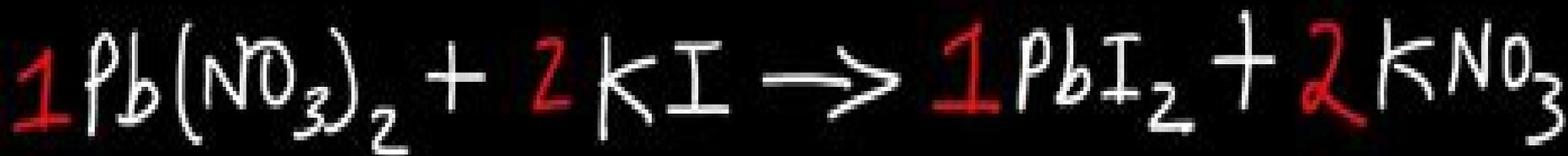


# Solution Stoichiometry



$$\frac{0.29 \text{ mol KI}}{\cancel{L}} \times \frac{0.037 \cancel{L}}{1} \times \frac{1 \text{ mol Pb}(\text{NO}_3)_2}{2 \cancel{\text{mol KI}}} \times \frac{1}{0.042 \text{ L}}$$

$$= 0.1277 \text{ M}$$

$$M = \frac{\text{mol}}{\text{L}}$$

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**S Marginson**



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