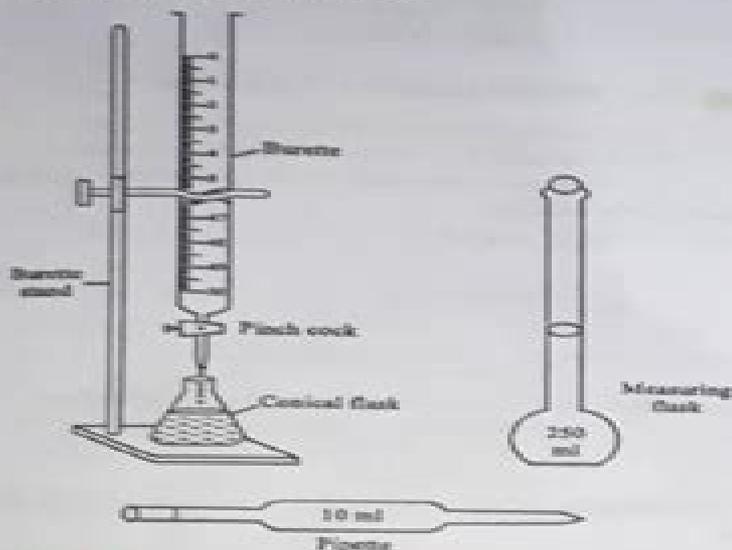


## EXPERIMENT NO. 8-1

**Object :** Prepare 250 ml standard solution of N/20 Ferrous ammonium sulphate and find out the strength in gram/litre of Potassium permanganate solution of unknown concentration by oxidation-reduction (Redox) single titration.

**Required materials :** Ferrous Ammonium Sulphate, Potassium Permanganate,  $H_2SO_4$ , Burette, Pipette, Burette stand, Chemical balance, Weight box, Funnel, Test-tube, Measuring flask, Watch glass, Marble tile, Conical flask etc.

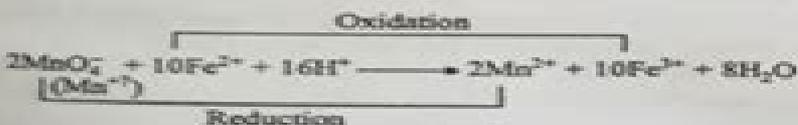


**Principle :** Acidic  $KMnO_4$  acts as a strong oxidising agent and Ferrous ammonium sulphate (Mohr's salt) acts as a strong reducing agent. When  $KMnO_4$  solution is added to a definite volume of Ferrous ammonium sulphate in the presence of dilute  $H_2SO_4$ , then Ferrous sulphate oxidises to Ferric sulphate. A light and stable pink colour is produced in the solution at the end point. Here  $KMnO_4$  acts as a self indicator.

**Chemical Equation :**



In Ionic form :



**Procedure :** (I) Preparation of standard solution of Ferrous ammonium sulphate of N/20 concentration :

Required quantity of FAS (W) =

$$\frac{\text{Equivalent mass (E)} \times \text{Normality (N)} \times \text{Volume (V)}}{1000}$$

Equivalent mass of FAS (E) = 392 gm,

Normality of FAS (N) =  $\frac{1}{20}$

Volume of FAS (V) = 250 ml.

$$\text{Required quantity of FAS (W)} = \frac{392 \times \frac{1}{20} \times 250}{1000} = 4.9 \text{ gm.}$$

**Observation table :**

S. No.	Reading of Pipette (ml)	Reading of Burette		Volume of used $KMnO_4$ (ml)	Concordant volume of used $KMnO_4$
		Initial reading	Final reading		
1.	10 ml	0.0 ml	9.5 ml	9.7 ml	9.5 ml
2.	10 ml	9.5 ml	19.2 ml	9.5 ml	
3.	10 ml	19.2 ml	28.7 ml	9.5 ml	

**Calculation :** According to the law of equivalence :

$$N_1 V_1 = N_2 V_2$$

Where,  $N_1$  = Normality of known FAS solution =  $\frac{1}{20}$

$V_1$  = Volume of known FAS solution = 10 ml

$N_2$  = Normality of unknown  $KMnO_4$  solution = ?

$V_2$  = Concordant volume of unknown  $KMnO_4$  = 9.5 ml.

By putting the values in,

$$N_1 V_1 = N_2 V_2$$

$$= \frac{1}{20} \times 10 = N_2 \times 9.5$$

$$N_2 = \frac{10}{20 \times 9.5} = 0.052 \text{ N.}$$

Concentration of unknown  $KMnO_4$  solution in gram/litre

= Normality ( $N_2$ ) of  $KMnO_4$  × Equivalent mass of  $KMnO_4$

= 0.052 × 31.6

= 1.64 gm/litre.

Weigh exactly 4.9 gm Ferrous ammonium sulphate in a clean and dry watch glass and transfer it to a 250 ml capacity measuring flask by a dry funnel, in which a little distilled water is already taken, flask is shaken to dissolve the entire crystalline substance in water, level of water is made upto the mark. The solution should be homogenous.

(II) Titration : Firstly burette is cleaned by rinsing it with water and then with  $KMnO_4$  solution. Then unknown  $KMnO_4$  solution is filled in the burette upto the zero mark and initial reading is noted. A clean pipette is rinsed with FAS solution once, then 10 ml of standard FAS solution is taken in this pipette and poured into a clean conical flask. 5 ml of dilute  $H_2SO_4$  is mixed in it. Conical

# Xi Chemistry Practical Procedure Volumetric Analysis

**Christian Drosten**



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