

10.1 The mole represents a large number of extremely small particles.

10.1 Measuring Matter

10.1.1 Chemists use the mole to count atoms, molecules, ions, and formula units.

10.2 Mass and the Mole

10.2.1 A mole always contains the same number of particles, however, moles of different substances have different masses.

10.3 Molar Mass

10.3.1 The molar mass of a compound can be calculated from its chemical formula and can be used to convert from mass to moles of that compound.

10.4 Empirical and Molecular Formulas

10.4.1 A molecular formula of a compound is a whole-number multiple of its empirical formula.

10.5 Formulas of Hydrates

10.5.1 Hydrates are ionic compounds in which water molecules are trapped.

ChemFacts

- The U.S. Mint produces annually 2.5 billion pennies, which are the most produced items in the world.
- The average daily intake of water is 2.5 liters, and the average person consumes 2.5 liters of water each day.
- The average person consumes 2.5 liters of water each day.



Chapter 10 The Mole Middlesex County Vocational And

Sabine Zange



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