

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the parentheses.

- 1.** mole: gram-molecular weight
- 2.** empirical formula: simplest formula
- 3.** molar mass: weight of one mole of a substance
- 4.** 10^{-24} : unit of mass
- 5.** a description of the atoms present in a substance: composition
- 6.** the lowest whole number ratio of the atoms of the elements in a compound: empirical formula
- 7.** formula weight: weight of one mole of a substance
- 8.** the number of moles of a substance: molar mass
- 9.** the number representing 6.02×10^{23} elementary particles of a substance: Avogadro's number
- 10.** 10^{-24} particles: atomic mass unit
- 11.** the number of elementary particles of a substance: Avogadro's number

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Completion Use the completion method to check your knowledge of the terms and facts concerning the concepts involved in the section. Each blank space represents a term, fact or phrase, as written.

Knowing what mole is, counting, weighing, and volume for a single quantity called the mole. Avogadro's number particles are there in a substance is Avogadro's number.

In the case of a compound, atoms are combined together the molar mass of the substance of the compound. When we substitute the unit mass for atoms (the atomic mass) of the compound, there are Avogadro's number particles the mole of a substance.

True or False

10.1 True A mole is a unit of mass.

10.2 True Avogadro's number is 6.02×10^{23} .

10.3 True The elementary particles of a substance is the Avogadro's number.

10.4 True The mole mass of a compound of one mole of sodium and two atoms of chlorine.

10.5 True The mole mass of nitrogen gas is 28.02 g.

Problems

Chapter 10 Chemical Quantities D Practice Answers

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