

Stoichiometric Calculations

Mole to Mole Calculations

1. The octane present in gasoline burns according to the following equation.



- (a) How many moles of O_2 are needed to react fully with 7 moles of octane (C_8H_{18})?
- (b) How many moles of CO_2 can form from 5 mole of octane?
- (c) If the reaction is to be used to synthesize 8 moles of CO_2 , how many moles of oxygen are needed?
- (d) How many moles of carbon dioxide are formed if 17 moles of oxygen are reacted?
2. Butane combines with oxygen in a combustion reaction to produce carbon dioxide and water.
- (a) Write the balanced equation for this reaction.
- (b) If 37 moles of water are produced, how many moles of butane were reacted?
- (c) 29 moles of oxygen are reacted with the butane, how many moles of carbon dioxide are produced?
- (d) How many moles of butane are required to react with 17.5 moles of oxygen?
3. Consider the following reaction:



- (a) How many moles of NO are produced by the reaction of 4.0 moles of copper with HNO_3 ?
- (b) How many moles of HNO_3 are required to react completely with 5.0 mol of copper?
- (c) What mass of NO is produced when 6.35 g of copper is reacted with HNO_3 ?

Chapter 122 Stoichiometric Calculations

R Barnett



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