CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

Solutions

Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.
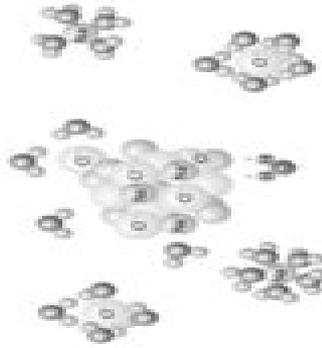
Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solute	solvent

Air is (1) __solution__ of oxygen gas dissolved in nitrogen
gas. The oxygen in air is the (2) __solute__ and nitrogen is the
(3) __solvent__. Because oxygen gas dissolves in a solvent, oxygen gas
is (4) __soluble__. A substance A substance that does not dissolve is
(5) __insoluble__. (6) __Liquid__ solutions are the most common
type of solutions. If one liquid is soluble in another liquid, such as water, we say the
two liquids are (7) __miscible__. However, if one liquid is insoluble in another,
the liquids are (8) __immiscible__.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.

9. If a solute is in solution in which the solvent is water, What is solvation?
__Solvation is the process of surrounding solute particles with solvent particles to form a solution.__

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?
__The sodium and chloride ions are separated and surrounded by the water molecules.__

11. Explain the orientation of the water molecules around the sodium ions and chloride ions.
__Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule. Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.__

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?
__The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.__

13. List three ways that the rate of solvation may be increased.
__stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent__

In your textbook, read about heat of solution, solubility and factors that affect solubility.

For each statement below, write true or false.

__true__ 14. The overall energy change that occurs when a solution forms is called the heat of solution.

__false__ 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

__false__ 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

__true__ 17. In a saturated solution, solvation and crystallization are at equilibrium.

__true__ 18. Additional solute can be dissolved in an unsaturated solution.

__true__ 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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M. C. Gore

Chapter 18 Study Guide For Content Mastery Answer Key:

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