

Chapter 8

Solutions, Acids, and Bases

Section 8.1 Formation of Solutions

(pages 128–133)

Three factors determine the extent of a solution: the processes that occur when components interact; whether the properties of a mixture compare with those of its individual solute.

Properties of Liquid Solutions (page 128)

1. **Thermodynamic properties of a solution can differ from those of its components.**
 - a. _____, _____, _____
 - b. _____, _____, _____
 - c. _____, _____, _____
2. **Comparative volatility of two substances is not additive.**
_____ and _____ differ due to a joint involvement of physical forces and concept of a good solvent-solute bond.
3. **Order, like structure, is relative, not dependent on what is fixed.**
 - a. The most intense forces are hydrogen.
 - b. The next most intense forces are dipole-dipole.
 - c. The next most intense are London forces.
 - d. The next most intense are London forces.

Heat of Solution (page 128)

1. **Calculating solution enthalpy for a solid salt:** _____, _____
process is exothermic heat.
2. **Calculating solution enthalpy for a gas:** _____, _____
process is exothermic heat.
3. **In the following answers, use an ionic hydrating ion attraction energy table, particle size, and the distances among solute particles. Use**
 - a. **Relative heat of the process.**

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J. G. R. Briggs



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