

9/10

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Calorimetry and Hess's Law

Prelab

1. Zinc metal reacts with hydrochloric acid according to the balanced equation:



When 0.103 g of Zn (s) is combined with enough HCl to make ^{50.0 g} 50.0 mL of solution in a coffee-cup calorimeter, all of the zinc reacts, raising the temperature of the solution from 22.5°C to 23.7°C. Find ΔH_{rxn} for this reaction as written. Use 1.0 g/mL for the density of the solution and 4.18 J/(g·°C) as the specific heat capacity. Ignore the heat absorbed by the calorimeter.

$$\Delta T = 23.7^\circ\text{C} - 22.5^\circ\text{C} = 1.2^\circ\text{C}$$

$$q_{\text{rxn}} = - (m_w)(c_w)(\Delta T_w)$$

$$q_{\text{rxn}} = - (50.0 \text{ g})(4.18 \text{ J/g}^\circ\text{C})(1.2^\circ\text{C})$$

$$q_{\text{rxn}} = -250.8 \text{ J} = -0.2508 \text{ kJ}$$

$$0.103 \text{ g Zn} \times \frac{1 \text{ mol Zn}}{65.38 \text{ g/mol Zn}} = 0.0015754 \text{ mol Zn}$$

$$\Delta H_{\text{rxn}} = \frac{-0.2508 \text{ kJ}}{0.0015754 \text{ mol Zn}}$$

$$\Delta H_{\text{rxn}} = 159.2 \text{ kJ/mol}$$

2. Calculate ΔH_{rxn} for the reaction:



Use the following reactions and given ΔH 's.



x2

$$\Delta H = -74.6 \text{ kJ}$$

$$\Delta H = -95.7 \text{ kJ}$$

$$\Delta H = -92.3 \text{ kJ} \times 2$$

$$-369.3$$

$$\Delta H_{\text{rxn}} = +74.6 + -95.7 + -369.3$$

$$\Delta H_{\text{rxn}} = 390.3$$

Chem Hess Law Lab Answer

David Kirk



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