

CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY1) Which of the following are weak electrolytes? HCl, $\text{HC}_2\text{H}_3\text{O}_2$, NH_3 , KCl

A) HCl, KCl

B) HCl, $\text{HC}_2\text{H}_3\text{O}_2$, NH_3 , KClC) $\text{HC}_2\text{H}_3\text{O}_2$, KClD) $\text{HC}_2\text{H}_3\text{O}_2$, NH_3 E) HCl, $\text{HC}_2\text{H}_3\text{O}_2$, KCl*Answer: D Diff: 2*2) What are the spectator ions in the reaction between KOH (aq) and HNO_3 (aq)?A) K^+ and H^+ C) K^+ and NO_3^- E) OH^- onlyB) H^+ and OH^- D) H^+ and NO_3^- *Answer: C Diff: 2*3) The balanced net ionic equation for precipitation of CaCO_3 when aqueous solutions of Na_2CO_3 and CaCl_2 are mixed is _____.A) $2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Na}_2\text{CO}_3(\text{aq})$

E)

B) $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq})$ $\text{Na}_2\text{CO}_3(\text{aq}) + \text{CaCl}_2(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{CaCO}_3(\text{s})$ C) $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{NaCl}(\text{aq})$ D) $\text{Ca}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CaCO}_3(\text{s})$ *Answer: D Diff: 3*4) Which of the following are strong acids? HI, HNO_3 , HF, HBr

A) HF, HBr

C) HI, HF, HBr

E) HI, HNO_3 , HBrB) HI, HNO_3 , HF, HBrD) HNO_3 , HF, HBr*Answer: E Diff: 2*

5) A neutralization reaction between an acid and a metal hydroxide produces _____.

A) water and a salt

C) oxygen gas

E) ammonia

B) hydrogen gas

D) sodium hydroxide

*Answer: A Diff: 2*6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na

B) Au

C) Fe

D) Ca

E) Ag

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?



A) Ni

B) Ag

C) Cu

D) Pb

E) N

Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100

B) 0.0500

C) 0.0333

D) 0.0250

E) 125

Answer: A Diff: 3

Chemistry Chapter 8 Test

Richard H. Langley, John Moore



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