

Chapter Test II

A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

Column I

- _____ 1. An unbalanced chemical equation for a reaction that obeys the law of conservation of matter.
- _____ 2. The number of protons in the nucleus of an atom of an element.
- _____ 3. The number of electrons in a neutral atom.
- _____ 4. The ratio of the actual weight of an element to the standard weight of 16.
- _____ 5. The number of protons and neutrons in the nucleus of an atom.
- _____ 6. The number of protons and neutrons in the nucleus of an atom.

Column II

- a. atomic weight
- b. limiting reagent
- c. molecular weight
- d. molar weight
- e. atomic number
- f. mass number

B. Multiple Choice

Choose the correct answer for each question.

- _____ 1. How many molecules of H_2O are formed when 100 g of H_2 reacts with 100 g of O_2 ?

<ol style="list-style-type: none"> a. 1000 molecules b. 10000 molecules c. 100000 molecules d. 1000000 molecules 	<ol style="list-style-type: none"> e. 10000000 molecules f. 100000000 molecules g. 1000000000 molecules h. 10000000000 molecules
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- _____ 2. Which of the following is not a chemical change?
 - a. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
 - b. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
 - c. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$
 - d. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2$
- _____ 3. The ratio of the actual weight of an element to the standard weight of 16 is called

<ol style="list-style-type: none"> a. atomic weight b. molecular weight 	<ol style="list-style-type: none"> c. atomic number d. mass number
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Chemistry Chapter 12 Test B Stoichiometry

John Moore, Richard H. Langley



Chemistry Chapter 12 Test B Stoichiometry:

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