

Chemistry - Chapter 7 Study Guide

Section 7.1:

- *Valence electrons* – electrons in the outer energy level
 - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
 - Look at the group number
 - Group 1A = 1 valence electron
 - Group 7A = 7 valence electrons
 - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
 - Written using the following steps:
 1. Write the element symbol – representing the nucleus and inner electrons
 2. Put one dot for each valence electrons,
 3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
 - Lose electrons to attain a noble gas configuration
 - Written like Na^{1+}
 - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
 - Gain electrons to attain a noble gas configuration
 - Written like P^{3-}
 - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
 - Also called salts
 - Made from a cation and an anion
 - Which means they're made of a metal and a non-metal
- ***if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
 - Written like Ca_3P_2

Chemistry Chapter 7 Study Guide For Content Mastery Answers

Kathleen Armour



Chemistry Chapter 7 Study Guide For Content Mastery Answers:

Chemistry McGraw-Hill Staff, 2001-07 **For All Practical Purposes** Consortium for Mathematics and Its Applications (U.S.), 2003 The sixth edition of the acclaimed classroom favorite offer a number of new features to help instructors strengthen the mathematical literacy of their students **Organic Chemistry** Francis A. Carey, 2000 A Market Leading Traditional Approach to Organic Chemistry Throughout all seven editions Organic Chemistry has been designed to meet the needs of the mainstream two semester undergraduate organic chemistry course This best selling text gives students a solid understanding of organic chemistry by stressing how fundamental reaction mechanisms function and reactions occur With the addition of handwritten solutions new cutting edge molecular illustrations updated spectroscopy coverage seamless integration of molecular modeling exercises and state of the art multimedia tools the 7th edition of Organic Chemistry clearly offers the most up to date approach to the study of organic chemistry **Prentice Hall Chemistry**, 2000 **Chemistry** William L. Masterton, 1996-08 This new edition of CHEMISTRY PRINCIPLES AND REACTIONS continues to provide students with the core material essential to understanding the principles of general chemistry Masterton and Hurley cover the basics without sacrificing the essentials appealing to several markets Appropriate for either a one or two semester course CHEMISTRY PRINCIPLES AND REACTIONS Fifth Edition is three hundred pages shorter than most general chemistry texts and lives up to its long standing reputation as THE student oriented text Though this text is shorter in length than most other General Chemistry books it is not lower in level and with the addition of the large volume of content provided by the revolutionary GENERAL CHEMISTRY INTERACTIVE 3 0 CD ROM that is included with every copy it has a depth and breadth rivaling much longer books **DAT** Aftab S. Hassan, 1999 At Lippincott Williams and Wilkins our objective is to help busy students and actual practitioners prepare for the wide array of exams required in the health sciences Our authors and editors are skilled question writers and all our products provide the practice material that students require That s why all of our review series products the top preparatory guides formerly known as the BETZ Guides NMS BRS Recall and High Yield are the first choices of students at schools throughout the country At Lippincott Williams and Wilkins we want to help you master the science of review **Resources in Education**, 1988 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Chemistry of Matter**, 1993 As the reader reads this textbook he will learn about the interactions of matter that can occur in a test tube in nature and even inside himself P 9 *Essentials of Ecology* George Tyler Miller, 2006-09 G Tyler Miller s worldwide bestsellers have evolved right along with the changing needs of your diverse student population Focused specifically on energizing and engaging all your students Miller and new contributor Scott Spoolman have been at work scrutinizing every line enhancing clarifying and streamlining to reduce word density as well as updating with the very latest environmental news and research The resulting texts are shorter clearer and so engaging that your students will actually want to read their assignments The ideal alternative to ecology texts that tend to be too difficult

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