

Qualitative Analysis Guide Sheet for C.S.E.E. Chemistry Practical

Systematic Qualitative Analysis can be grouped into three stages:

- 1) Preliminary tests
- 2) Tests in solution
- 3) Confirmation tests

The ions expected in the scheme include:

Cations: NH_4^+ Ca^{2+} Fe^{3+} Cu^{2+} Zn^{2+} Pb^{2+} Na^+ Fe^{2+}

Anions: CO_3^{2-} HCO_3^- NO_3^- SO_4^{2-} Cl^-

Recording experimental results

The recorded results should appear in a tabular form where three columns showing the brief explanation of the experimental procedure, observations made, and inferences must appear.

The report should be in the following table:

TEST	OBSERVATIONS	INFERENCE

OBSERVATIONS:

When doing experiments, leave the chemicals to stand for a few minutes as some changes may take a little time to be seen. For example, most precipitates form instantly, but others take a minute or two.

Thus observations are generally changes which can be seen when a chemical substance is subjected to an experiment (test). These include:

- a) colour change
- b) precipitates formed
- c) gelatinous changes of liquid substances
- d) gases evolved
- e) light (noise) produced

INFERENCE:

Inferences are essential in experimental procedure as they explain about the identity of the unknown chemical substances. Usually inferences must be specific. For example a common mistake of making deductions is to write "sodium, Na^+ ". Does this mean sodium Na^+ ions are present or absent? Students should remember to write specific inferences. Students should also remember that not all experiments will give a visible change. Sometimes if there is no change, it can help to make an inference about the absence of certain ions. For example, if a barium nitrate solution added to an unknown substance and there is no reaction, this implies that sulphate ions are absent because the presence of sulphate ions is shown by formation of white precipitate.

CONCLUSION:

Before proceeding to the next test, observations and inferences of the earlier experiment must be written. After all experiments you may be asked to make conclusions about the chemical ions or nature of the chemical substances found in the samples. This can be obtained by combining together all inferences made earlier in the successive tests.

Chemistry Practical Qualitative Analysis Sheet

Purdue University



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Practical Chemistry. The Principles of Qualitative Analysis Sir William Augustus Tilden, 1880 **A Laboratory Manual of Organic Chemistry for Beginners** Arnold Frederick Holleman, 1913 **Practical Chemistry; the Principles of Qualitative Analysis** Sir William Augustus Tilden, 2013-09 This historic book may have numerous typos and missing text Purchasers can usually download a free scanned copy of the original book without typos from the publisher Not indexed Not illustrated 1886 edition Excerpt PRACTICAL CHEMISTRY PART I PRELIMINARY To make a Washing Bottle Take a clean widemouth bottle holding about a pint and fit a good sound cork tightly into the neck Then take a piece of stout glass tubing about inch external diameter and with the aid of a cork borer bore two holes through the cork parallel to each other and of such a size that the tube will pass through with some friction If you have not a cork borer of the exact size required take one a little smaller and then enlarge the hole by means of a round file Cut off a piece of the tube of about twice the height of the bottle Draw off one end to a point in the blowpipe flame Cut it off with a file and make the orifice smooth by holding it in the flame for a minute or two taking care not to close it completely Then in a common fish tail gas flame bend the tube so as to B form an angle of about 60 at such a distance from the wide end that when the tube is passed through the cork it will reach well to the bottom of the bottle Into the second hole in the cork another piece of the same tubing is fitted but this is cut off just below the cork and is bent to form an angle of about 120 The open end projecting upwards should be made smooth at the edge by holding it for a few minutes in quantities of water into test tubes and for general purposes as a store of water It is used by blowing gently into the upturned end of the mouthpiece and so producing a pressure upon the surface of the water in the bottle To Fix a Platinum Wire into a Glass Holder Take six inches of narrow glass tubing and by holding it in the Bunsen flame soften it so that it may be drawn out in the middle When cold scratch it at the middle with a file and break it

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