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SAP BusinessObjects Web Intelligence User's Guide

Business Objects Web Intelligence User Guide

D Kirk



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web limiting reactant theoretical and percentage yields key $4 KO_2 + 2 H_2O \rightarrow 4 KOH + O_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{CuNO}_3 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$

 8.3 mole H_2S x 18g H_2O / 1 mole H_2O = 149.4g H_2O is your theoretical yield so percent yield = $\frac{137.1\text{g } \text{H}_2\text{O}}{149.4\text{g } \text{H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mole}} = 0.666 \text{ moles } \text{O}_2$

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web a write a balanced equation for the reaction

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

 b determine the theoretical yield of Al_2O_3
 $\frac{160 \text{ g } \text{Al}}{27 \text{ g/mol}} = 5.93 \text{ mol } \text{Al}$
 $\frac{5.93 \text{ mol } \text{Al}}{4} \times 2 = 2.96 \text{ mol } \text{Al}_2\text{O}_3$
 $2.96 \text{ mol } \text{Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 302 \text{ g } \text{Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 = \frac{260 \text{ g}}{302 \text{ g}} \times 100 = 86.1\%$

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web 7. Na₂CO₄ is the limiting reactant percent yield = 86.68% only four molecules can be made 9. this amount cannot be weighted by ordinary balances and is worthless 10. nitrogen is the limiting reagent 11. yes methane is the limiting reagent 12. C is the limiting reagent 4.33 g of H₂ are left over 13.

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