

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the vocabulary list.

- | | |
|-------------------------|---------------------------|
| _____ mole | _____ percent composition |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ molar mass |
| _____ molar mass | _____ molar mass |
10. _____ is a description of the relative amounts of each element in a compound. _____ percent composition
11. _____ is the lowest whole number ratio of the atoms of the elements in a compound. _____ empirical formula
12. _____ is the mass of one mole of any gas at STP. _____ 22.4 L
13. _____ is the mass of one mole of a compound. _____ molar mass
14. _____ is the amount corresponding to 6.02 x 10²³ representative particles of a substance. _____ mole
15. _____ is 6.02 x 10²³ particles. _____ Avogadro's number
16. _____ is the mass of one mole of a substance at STP. _____ molar mass
17. _____ is the mass of one mole of a gas at STP. _____ molar mass
18. _____ is the mass of one mole of a gas at STP. _____ molar mass

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Complete the following activities to check your knowledge of the terms and your understanding of the concepts involved with the terms. Each activity has a corresponding answer sheet given as well.

1. _____ is the amount corresponding to 6.02 x 10²³ representative particles of a substance. _____ Avogadro's number

2. _____ is the mass of one mole of a compound, element, and together the _____ molar mass. _____ of the atoms making up the compound. _____ is the mass of one mole of a gas at STP. _____ is the mass of one mole of a gas at STP. _____ is the mass of one mole of a gas at STP.

Practice Problems

1. _____ is the mass of one mole of a substance. _____ molar mass

2. _____ is the mass of one mole of a gas at STP. _____ molar mass

3. _____ is the mass of one mole of a gas at STP. _____ molar mass

4. _____ is the mass of one mole of a gas at STP. _____ molar mass

5. _____ is the mass of one mole of a gas at STP. _____ molar mass

6. _____ is the mass of one mole of a gas at STP. _____ molar mass

Answers

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David Kirk



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