

Chapter 12 Study Guide - Stoichiometry

TEST: Thursday 3/10/11

12.1 The Arithmetic of Equations

1. Define *stoichiometry*

2. What is always conserved in every chemical reaction?

12.2 Chemical Calculations

1. Define *mole ratio*

2. In a balanced chemical equation, where can one find the numbers to make up the mole ratio?

3. For the following reaction: $2\text{H}_2\text{S (g)} + 3\text{O}_2 \text{(g)} \rightarrow 2\text{SO}_2 \text{(g)} + 2\text{H}_2\text{O (g)}$
 - a. What is the mole ratio of hydrogen sulfide to water?

 - b. What is the mole ratio of oxygen to sulfur dioxide

 - c. What is the mole ratio of water to oxygen?

Chapter 12 Study Guide For Content Mastery Stoichiometry

Georg G. Behrens



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