

## Calculating Specific Heat

An iron skillet has a mass of 500.0 grams. The specific heat of iron is  $0.449 \text{ J/g}^\circ\text{C}$ . How much heat must be absorbed to raise the skillet's temperature by  $95.0^\circ\text{C}$ ?

### 1 Read and Understand

*What information are you given?*

Mass of iron,  $m = 500.0 \text{ g}$

Specific heat of iron,  $c = 0.449 \text{ J/g}^\circ\text{C}$

Temperature change,  $\Delta T = 95.0^\circ\text{C}$

### 2 Plan and Solve

*What unknown are you trying to calculate?*

Amount of heat needed,  $Q = ?$

*What formula contains the given quantities and the unknown?*

$$Q = m \times c \times \Delta T$$



*Replace each variable with its known value.*

$$\begin{aligned} Q &= 500.0 \text{ g} \times 0.449 \text{ J/g}^\circ\text{C} \times 95.0^\circ\text{C} \\ &= 21,375 \text{ J} = 21.4 \text{ kJ} \end{aligned}$$

### 3 Look Back and Check

*Is your answer reasonable?*

Round off the data to give a quick estimate.

$$Q = 500 \text{ g} \times 0.5 \text{ J/g}^\circ\text{C} \times 100^\circ\text{C} = 25 \text{ kJ}$$

This is close to 21.4 kJ, so the answer is reasonable.

Math Practice #  
1- 5 Page 477

Chapter 16 Thermal Energy And Heat Math Practice  
Page 477

**KJ Lindholm-Leary**



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### **Table of Contents Chapter 16 Thermal Energy And Heat Math Practice Page 477**

1. Understanding the eBook Chapter 16 Thermal Energy And Heat Math Practice Page 477
  - The Rise of Digital Reading Chapter 16 Thermal Energy And Heat Math Practice Page 477
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 16 Thermal Energy And Heat Math Practice Page 477
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 16 Thermal Energy And Heat Math Practice Page 477
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 16 Thermal Energy And Heat Math Practice Page 477
  - Personalized Recommendations
  - Chapter 16 Thermal Energy And Heat Math Practice Page 477 User Reviews and Ratings
  - Chapter 16 Thermal Energy And Heat Math Practice Page 477 and Bestseller Lists
5. Accessing Chapter 16 Thermal Energy And Heat Math Practice Page 477 Free and Paid eBooks
  - Chapter 16 Thermal Energy And Heat Math Practice Page 477 Public Domain eBooks
  - Chapter 16 Thermal Energy And Heat Math Practice Page 477 eBook Subscription Services
  - Chapter 16 Thermal Energy And Heat Math Practice Page 477 Budget-Friendly Options
6. Navigating Chapter 16 Thermal Energy And Heat Math Practice Page 477 eBook Formats

- ePub, PDF, MOBI, and More
  - Chapter 16 Thermal Energy And Heat Math Practice Page 477 Compatibility with Devices
  - Chapter 16 Thermal Energy And Heat Math Practice Page 477 Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chapter 16 Thermal Energy And Heat Math Practice Page 477
    - Highlighting and Note-Taking Chapter 16 Thermal Energy And Heat Math Practice Page 477
    - Interactive Elements Chapter 16 Thermal Energy And Heat Math Practice Page 477
  8. Staying Engaged with Chapter 16 Thermal Energy And Heat Math Practice Page 477
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chapter 16 Thermal Energy And Heat Math Practice Page 477
  9. Balancing eBooks and Physical Books Chapter 16 Thermal Energy And Heat Math Practice Page 477
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chapter 16 Thermal Energy And Heat Math Practice Page 477
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chapter 16 Thermal Energy And Heat Math Practice Page 477
    - Setting Reading Goals Chapter 16 Thermal Energy And Heat Math Practice Page 477
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chapter 16 Thermal Energy And Heat Math Practice Page 477
    - Fact-Checking eBook Content of Chapter 16 Thermal Energy And Heat Math Practice Page 477
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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