

14 • The Chemistry of Acids & Bases**STUDY QUESTIONS & PROBLEMS - KEY**

1. For the following aqueous equilibria, designate the Bronsted-Lowry conjugate acid-base pairs and establish the weaker side:



2. Write the name and formula for the conjugate bases of the following:



3. Complete the Bronsted-Lowry equilibria, label the components acid or base, and pair up the conjugate acid-base pairs:



4. Is the monohydrogen phosphate ion HPO_4^{2-} amphiprotic? -- YES

If so, write the formulas of its conjugate acid and its conjugate base. CA = H_2PO_4^- and CB = PO_4^{3-}

Chapter 17 Study Acids Bases

Rachel S Tattersall



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