

# Chapter 17 Thermochemistry Study Guide

## 17.1 – 17.2 Thermochemical Equations

*General Chemistry*

1. Make the following conversions:
  - a. 444 cal to joules =  $1.86 \times 10^3$  J
  - b. 1.8 kJ to joules =  $1.8 \times 10^3$  J
  - c. 0.45 kJ to calories =  $1.1 \times 10^2$  cal
2. Classify each of these processes as endothermic or exothermic:
  - a. condensing steam – **exo**
  - b. burning alcohol – **exo**
  - c. evaporating alcohol – **endo**
  - d. baking a potato – **endo**
3. The specific heat capacity for silver is  $0.24 \text{ J/g}^\circ\text{C}$ . Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(197.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

4. It takes 585 J of energy to raise the temperature of 125.6 g Hg from  $20.0^\circ\text{C}$  to  $53.5^\circ\text{C}$ . Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

5. A 46.2-g sample of copper is heated to  $95.4^\circ\text{C}$  and then placed in a calorimeter containing 75.0 g water at  $19.6^\circ\text{C}$ . The equilibrium temperature in the calorimeter is  $21.8^\circ\text{C}$ . Calculate the specific heat capacity of copper, assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

6. A 15.0-g sample of nickel metal is heated to  $100.0^\circ\text{C}$  and dropped into 55.0 g of water, initially at  $23.0^\circ\text{C}$ . Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is  $0.444 \text{ J/g}^\circ\text{C}$ .

$$\begin{aligned} -15.0(444)(T_f - 100.0) &= 55.0(4.184)(T_f - 23.0) \\ -6.66T_f + 666 &= 230.12T_f - 5292.76 \\ 5958.76 &= 236.78 T_f \\ &= T_f \\ &= 25.2^\circ\text{C} \end{aligned}$$

7. Chloe was running bath water and realized it was too hot. If she has 20.0 L of water in the tub at  $95^\circ\text{C}$  and then adds 15.8 L of water at  $75^\circ\text{C}$ , what will the final temperature of the water be?

$$\begin{aligned} -20(T_f - 95) &= 15.8(T_f - 75) \\ -20T_f + 1900 &= 15.8 T_f - 1185 \\ 3085 &= 35.8 T_f \end{aligned}$$

# Chapter 17 Thermochemistry Interpreting Graphics

## Answers

**Scott C. Dulebohn**



## **Chapter 17 Thermochemistry Interpreting Graphics Answers:**

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