

Reaction Rates and Mechanisms Quiz

Name: _____

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Multiple Choice; Circle the most correct answer. (10K, 10T)

1. Within a reaction with 2 reactants and 2 products, the reaction rate can be measured with respect to:
A) all 4 substances B) 3 of the substances C) 2 of the substances
D) only 1 of the substances E) not enough information to determine

2. The minimum energy required for a reaction to take place is called the:
A) Activation energy B) Activated complex C) Heat of reaction D) Enthalpy

3. A temporary structure with a high potential energy that has bonds half formed and half broken is:
a) a homogeneous catalyst b) an activated complex c) radical
d) an elementary process e) a heterogeneous catalyst

4. Consider the following reactions at 25°C.



Which of the following statements best describes the relative rates of the two reactions?

A. I and II are both fast. C. I is faster than II
B. I and II are both slow. D. II is faster than I

5. What property would be appropriate to measure the reaction rate in the following reaction:



a. change in conductivity b. change in mass c. change in colour
d. change in pressure e. change in volume

6. The theoretical effect of an increase in the initial concentration of a reactant can be explained in terms of collision theory because it affects:

- I. the collision geometry involved in the reaction
- II. the total number of collisions that occur
- III. the fraction of collisions that are effective
- IV. the required activation energy for a reaction

A) both I and IV B) I, II and III are true C) both III and IV D) both II and III E) II only

7. At a certain temperature, 2.50 g Ca reacts completely in 30.0 seconds.
The rate of consumption of Ca is:



A) 0.00208 mol/min B) 0.0833 mol/min C) 0.125 mol/min D) 5.00 mol/min

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