

Student Exploration: Calorimetry Lab

(Note: Gizmos can be done even on your phones! Contact Goff or Molik if you need your log-in information)

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.



Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.

1. On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).
 - A. What was the **Final temperature** of the copper and the water? 35 °C _____
 - B. How much did the temperature of the copper change? It went down 65 °C _____
 - C. How much did the temperature of the water change? _____ it went up 5 °C _____
2. Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Activity A: Heat transfer	<u>Get the Gizmo ready:</u> • Click Reset (↺).	
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Question: What factors determine how heat energy transfers between objects?

1. **Predict:** In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g of 30.0 °C water.
 - A. How do you think increasing the water's mass would affect the final temperature?

Lab Activity Chemical Reaction Answer Key Calorimetry

Lei Shi



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Table of Contents Lab Activity Chemical Reaction Answer Key Calorimetry

1. Understanding the eBook Lab Activity Chemical Reaction Answer Key Calorimetry
 - The Rise of Digital Reading Lab Activity Chemical Reaction Answer Key Calorimetry
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Activity Chemical Reaction Answer Key Calorimetry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Activity Chemical Reaction Answer Key Calorimetry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Activity Chemical Reaction Answer Key Calorimetry

- Personalized Recommendations
 - Lab Activity Chemical Reaction Answer Key Calorimetry User Reviews and Ratings
 - Lab Activity Chemical Reaction Answer Key Calorimetry and Bestseller Lists
5. Accessing Lab Activity Chemical Reaction Answer Key Calorimetry Free and Paid eBooks
 - Lab Activity Chemical Reaction Answer Key Calorimetry Public Domain eBooks
 - Lab Activity Chemical Reaction Answer Key Calorimetry eBook Subscription Services
 - Lab Activity Chemical Reaction Answer Key Calorimetry Budget-Friendly Options
 6. Navigating Lab Activity Chemical Reaction Answer Key Calorimetry eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Activity Chemical Reaction Answer Key Calorimetry Compatibility with Devices
 - Lab Activity Chemical Reaction Answer Key Calorimetry Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Activity Chemical Reaction Answer Key Calorimetry
 - Highlighting and Note-Taking Lab Activity Chemical Reaction Answer Key Calorimetry
 - Interactive Elements Lab Activity Chemical Reaction Answer Key Calorimetry
 8. Staying Engaged with Lab Activity Chemical Reaction Answer Key Calorimetry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Activity Chemical Reaction Answer Key Calorimetry
 9. Balancing eBooks and Physical Books Lab Activity Chemical Reaction Answer Key Calorimetry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Activity Chemical Reaction Answer Key Calorimetry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Lab Activity Chemical Reaction Answer Key Calorimetry
 - Setting Reading Goals Lab Activity Chemical Reaction Answer Key Calorimetry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Lab Activity Chemical Reaction Answer Key Calorimetry

- Fact-Checking eBook Content of Lab Activity Chemical Reaction Answer Key Calorimetry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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