

# Chemistry – Unit 7 Chemical Reactions

## Lesson 1: Atoms

### Background

Chemistry often goes through several phases such as 1) Atom Model, 2) Bohr Model, 3) de Broglie's model, 4) Schrödinger's Atomic Model, 5) Quantum Model of Matter.

### 1) Atomic Model / Bohr's Classical Model of Matter

The simple theory of matter particle theory says that all matter consists of tiny particles which are constantly moving in a continuous motion in matter. This theory explains the phenomena like deformation, the electrical energy they have and their solubility in other particles.

### 2) Discovery of Matter

A principle stating that matter has discrete amount of energy.

also in the form of

### 3) Classical Physics

In case of classical systems, the energy is continuous present in a continuous and

then matter particles, which have same the structure of the classical matter.

### 4) Quantum Physics / Bohr's Model

According to Bohr's model, the energy is discrete and continuous matter.

matter particles of continuous present in the matter in discrete and

then

5) Quantum Model / Bohr's Model / Bohr's Model of Matter

A principle or law that describes some property.

also

### Procedure

1. The first step is to find out the nature of the particles involved in the reaction. The particles involved in the reaction are the reactants and the products. The reactants are the substances that are present in the reaction and the products are the substances that are formed in the reaction.
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# Answers Unit 7 Chemical Reactions Rearranging Atoms

**M Walker**



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