

Titration of Aspirin Tablets

In this lab, you will determine the percent purity of two commercially available aspirin tablets using an acid-base titration. In general, an acid and a base react to produce a salt and water by transferring a proton (H^+):



The active ingredient in aspirin, and the chemical for which aspirin is the common name, is acetylsalicylic acid. To determine the amount of aspirin (acetylsalicylic acid) in a sample, the precise volume and concentration of the NaOH, and the overall reaction, must be known. The NaOH serves as a secondary standard, because its concentration can change over time. To find the precise concentration of the NaOH, it must be titrated against a primary standard, an acid that dissolves completely in water, has a high molar mass, that remains pure upon standing, and is not hygroscopic (tending to attract water from the air). Because sodium hydroxide is hygroscopic, it draws water from its surroundings. This means one cannot simply weigh out a sample of sodium hydroxide, dissolve it in water, and determine the number of moles of sodium hydroxide present using the mass recorded, since any sample of sodium hydroxide is likely to be a mixture of sodium hydroxide and water. Thus, the most common way to determine the concentration of any sodium hydroxide solution is by titration. Determining the precise concentration of NaOH using a primary standard is called standardization. You will first standardize your NaOH solution, and then use it to analyze aspirin tablets for their aspirin content and purity.

An aspirin tablet may also include inactive ingredients that help produce a consistent product for consumers. Think of the purpose these ingredients serve when considering whether your percent purity is reasonable for a commercial aspirin tablet.

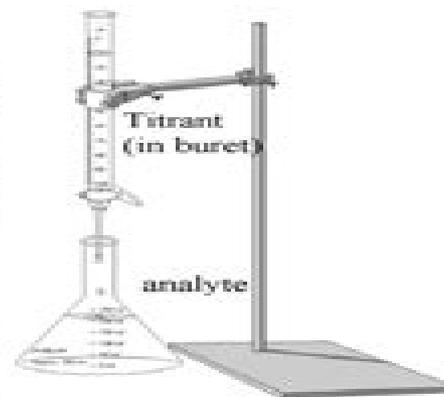
What is a titration?

A titration is a procedure for determining the concentration of a solution (the analyte) by allowing a carefully measured volume of this solution to react with another solution whose concentration is known (the titrant). The point in the titration where enough of the titrant has been added to react exactly with the analyte is called the equivalence point, and occurs when moles of titrant equals moles of analyte according to the balanced equation. For example, if a monoprotic acid (the analyte) is titrated with a strong base like sodium hydroxide (the titrant), the equivalence point occurs when



The equivalence point is often marked by an indicator, a substance that changes color at (or very near) the equivalence point.¹

There are many types of titrations. In this lab you will be performing an acid base titration.



¹ The point at which the indicator changes color is called the "endpoint". We will normally assume that the endpoint is equal to the equivalence point.

Analysis Of Aspirin Lab Report Conclusion

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