

Name \_\_\_\_\_

**Multiple Choice (10 points each)**

Identify the letter of the choice that best completes the statement or answers the question.

- E 1. For a certain process at 27°C,  $\Delta G = +210.6$  kJ and  $\Delta H = -168.2$  kJ. What is the entropy change for this process at this temperature? Express your answer in the form,  $\Delta S =$  \_\_\_\_\_ J/K.
- a. +628.3 J/K  
b. -141.3 J/K  
c. +141.3 J/K  
d.  $1.26 \times 10^3$  J/K  
e.  $-1.26 \times 10^3$  J/K

Answer:  $\Delta G = \Delta H - T\Delta S$   $T = 27^\circ\text{C} + 273 = 300\text{K}$

$$\Delta S = \frac{\Delta H - \Delta G}{T} = \frac{(-168.2\text{kJ}) - (+210.6\text{kJ})}{300\text{K}} = -\frac{1.26\text{kJ}}{\text{K}}$$

$$\left(\frac{-1.26\text{kJ}}{\text{K}}\right)\left(\frac{10^3\text{J}}{\text{kJ}}\right) = \boxed{\frac{-1.26 \times 10^3\text{J}}{\text{K}}}$$

- A 2. Which response includes all of the following substances that have  $\Delta H_f^\circ = 0$ , and no other substances?

- I. HCl(g)  
II. Na(s)  
III. HCl(aq)  
IV. F<sub>2</sub>(g)

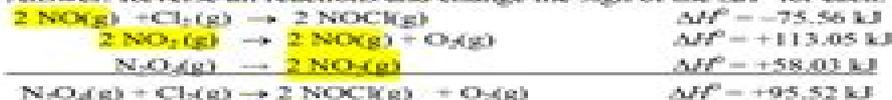
- a. II and IV  
b. II  
c. I and II  
d. I, II, III, and IV  
e. I, II, and IV

Answer: All elements in their natural standard state have a  $\Delta H_f^\circ = 0$ .

- A 3. From the following data at 25°C,

Calculate  $\Delta H^\circ$  at 25°C for the reaction below.

- a. +95.52 kJ  
b. +299 kJ  
c. -186.8 kJ  
d. -85.52 kJ  
e. -246.65 kJ

Answer: Reverse all reactions and change the sign of the  $\Delta H^\circ$  for each.

# Thermodynamics Problems And Answers

**R Barnett**



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