

1. List the intermolecular forces that can be found between molecules and define them.

Hydrogen bonding—attraction between the H bound to either F, O, or N, and a lone pair on one of those atoms.

Dipole-Dipole—attraction between the $-$ and $+$ ends of two molecules.

London Forces—small attraction that happens due to the presence of electrons.

2. For the following molecules, draw the Lewis structure, state whether it is polar/nonpolar, and list the intermolecular forces found in the molecules. Which would you expect to have the strongest intermolecular forces?

Water has the strongest intermolecular forces because it has the most hydrogen bonding, dipole-dipole, and London forces.

- a. H_2O
polar, H-bonding, Dip-dip, London



- d. OF_2
polar, Dip-Dip, London



- b. CH_4



Nonpolar, London

- e. NH_3
Polar, H-bonding, Dip-Dip, London



- c. CHCl_3
polar, Dip-Dip, London



- f. PH_3



Polar, Dip-Dip, London

Molecular Geometry And Intermolecular Forces Answer Key

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An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

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