

$$A_r(\text{Cu}) = 63.5 \quad A_r(\text{Ag}) = 108$$



$$y \qquad \qquad \qquad x$$

x = the quantity of deposited silver

y = the quantity of dissolved copper

$$\frac{63.5}{y} = \frac{2 \times 108}{x}$$

$$x - y = 101.52 - 100 \quad x = 1.52 + y$$

$$\frac{63.5}{y} = \frac{2 \times 108}{1.52 + x} \quad y = 0.63 \quad x = 2.15 \text{ g Ag}^+$$

Mass of silver nitrate:

$$\frac{840}{1000} \times 0.5 \times 170 = 71.4 \text{ g AgNO}_3$$

$$\frac{170 \text{ g AgNO}_3}{108 \text{ g Ag}} = \frac{71.4}{x} \quad x = 45.36 \text{ g Ag}^+$$

Silver consumed for participation

$$45.36 - 2.15 = 43.21 \text{ g Ag}^+$$

Total mass of chloride

$$\frac{108 \text{ g Ag}^+}{35.5 \text{ g Cl}^-} = \frac{43.2}{x} \quad x = 14.2 \text{ g Cl}^-$$

$$M_r(\text{NaCl}) = 58.5 \quad M_r(\text{KCl}) = 74.6$$

x = mass of NaCl in the mixture

y = mass of KCl in the mixture

$$\text{mass of Cl}^- \text{ in NaCl: } \frac{35.5 x}{58.5}$$

$$\text{mass of Cl}^- \text{ in KCl: } \frac{35.5 y}{74.6}$$

$$\frac{35.5 x}{58.5} + \frac{35.5 y}{74.6} = 14.2$$

$$x + y = 25$$

$$x = 17.6 \text{ g NaCl} \quad 70.4 \% \text{ NaCl}$$

$$y = 7.4 \text{ g KCl} \quad 29.6 \% \text{ KCl}$$

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Louis Munroe Dennis



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