

Analytical Chemistry  
3<sup>rd</sup> Long Exam

**PART I. MULTIPLE CHOICE.** Choose the letter that corresponds to the **BEST** answer. Write [ ] when the answer is **NOT** in the choices. (23 pts)

1. Transfer of electrons from one reactant to another

- a. precipitation rxn
- b. half rxn
- c. redox rxn
- d. neutralization rxn

2. What is the mole ratio of  $\text{Ce}(\text{IV})$  to  $\text{NO}_2^-$  for the oxidation of nitrite to nitrate using  $\text{Ce}^{4+}$  in the presence of an acid:



- a. 1:2
- b. 2:1
- c. 1:1
- d. 3:2

3. How many mole(s) of electron is(are) produced in the oxidation of sodium bismuthate ( $\text{NaBiO}_3$ ) to  $\text{BiO}^+$  and  $\text{H}_3\text{O}^+$  under acidic conditions?

- a. 1
- b. 2
- c. 3
- d. 4

4. Azide ( $\text{N}_3^-$ ) was added in the determination of the dissolved oxygen in water sample to remove interferences caused by the presence of what ion?

- a. chloride
- b. fluoride
- c. nitrite
- d. nitrate

5. Which of the following statements is(are) true?

- I. Prepared  $\text{KMnO}_4$  solutions should be stored in amber bottles
- II.  $\text{KMnO}_4$  solutions are boiled to remove  $\text{CO}_2$  which makes the solution basic
- III.  $\text{Mn}(\text{II})$  is the product of the degradation of  $\text{KMnO}_4$
- IV.  $\text{KMnO}_4$  solutions are filtered to remove manganese(IV) oxide

- a. I only
- b. I and II
- c. I and IV
- d. I, II and III

6. What is the oxidation state of carbon in oxalic acid?

- a. +2
- b. +4
- c. +6
- d. +7

7. The complex formed between  $\text{Cu}(\text{I})$  and 1,10-phenanthroline has a molar absorptivity of  $7000 \text{ L cm}^{-1} \text{ mol}^{-1}$  at 435 nm, the wavelength of maximum absorption. What is the absorbance of a  $5.17 \times 10^{-4} \text{ M}$  solution of the complex when measured in a 1.00-cm cell at 435 nm?

- a. 0.0431
- b. 0.0432
- c. 0.432
- d. 0.431

8. In the presence of iodine, thiosulfate ion is oxidized quantitatively to tetrathionate ion. What is the diatomic formula of tetrathionate ion?

- a.  $\text{S}_2\text{O}_7^{2-}$
- b.  $\text{S}_4\text{O}_6^{2-}$
- c.  $\text{S}_2\text{O}_8^{2-}$
- d.  $\text{S}_4\text{O}_8^{2-}$

9. In Winkler titration, the flocc is produced from

- a. oxidation of  $\text{I}^-$  to  $\text{I}_2$
- b. oxidation of  $\text{Mn}(\text{II})$  to  $\text{Mn}(\text{III})$
- c. reduction of  $\text{I}^-$  to  $\text{I}_2$
- d. reduction of  $\text{Mn}(\text{II})$  to  $\text{Mn}(\text{III})$

10. A component of a UV-Vis spectrophotometer that restricts the radiation being measured to a narrow band that is absorbed or emitted by the analyte.

- a. light source
- b. detector
- c. radiation filter
- d. wavelength selector

11. Component of a UV-Vis instrument that operates by blocking or absorbing all but a restricted band of radiation.

- a. light source
- b. detector
- c. radiation filter
- d. wavelength selector

12. Oxidation of  $\text{VO}^{2+}$  using permanganate or cerium as standard solution produces  $\text{V}(\text{OH})_5^{+2}$ , what is the oxidation state of vanadium in  $\text{V}(\text{OH})_5^{+2}$ ?

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