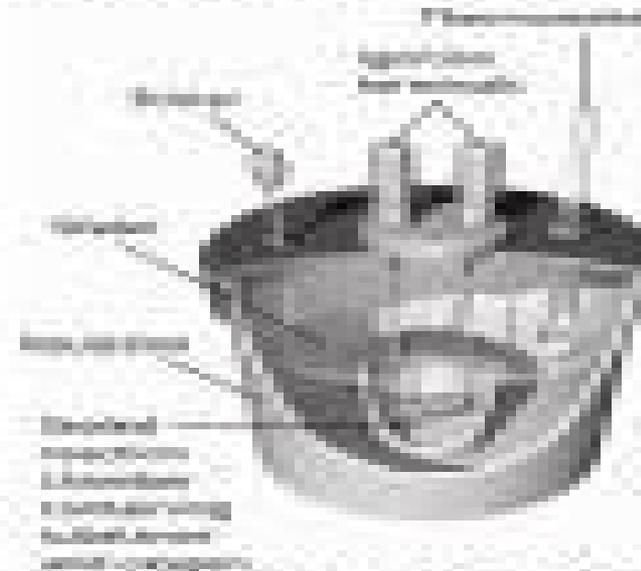


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- The calorimeter consists of a reaction chamber in which the reaction takes place, surrounded by water. The whole system is insulated so that heat is lost from the reaction chamber. The calorimeter is placed in the water, which is the surrounding substance.
- When the reaction takes place, heat is released or absorbed. This heat is transferred to the water, which is the surrounding substance. The temperature of the water is measured.
- When the reaction takes place, heat is released or absorbed. This heat is transferred to the water, which is the surrounding substance. The temperature of the water is measured.
- This is the calorimeter constant. It is used to calculate the heat of the reaction.
- When the reaction takes place, heat is released or absorbed. This heat is transferred to the water, which is the surrounding substance. The temperature of the water is measured.
- Calculate the amount of heat as the reaction takes place inside the chamber. The heat is transferred to the water, which is the surrounding substance. The temperature of the water is measured.
- Assuming that no heat is lost from the calorimeter, the amount of heat released or absorbed by the reaction is equal to the amount of heat gained or lost by the water.
- Since the power released from the reaction is equal to the amount of heat released or absorbed, the amount of heat released or absorbed is equal to the amount of heat gained or lost by the water.

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