

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 9

## STUDY GUIDE FOR CONTENT MASTERY

## Covalent Bonding

### Section 11 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	pi bond
---------------	----------	------------	------------	---------

When sharing electrons occurs, the attachment between atoms that results is called a(n) **covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **exothermic**. When two or more atoms bond by means of this type sharing, the resulting particle is called a(n) **molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **pi bond**.

In your textbook, read about single and multiple bonds and bond strength.

Circle the letter of the choice that best completes the statement or answers the question.

- Is a pair from two elements such as hydrogen, nitrogen, and oxygen normally covalent?
  - an ionic atom
  - a molecule containing three atoms
  - a molecule containing two atoms
  - a molecule containing four atoms
- How many electrons are shared in a double covalent bond?
  - none
  - two
  - four
  - eight
- Bond length is the distance between
  - two molecules of the same substance
  - the electrons in two attached atoms
  - the nuclei of two attached atoms
  - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
  - the shorter the bond, the stronger the bond
  - the shorter the bond, the weaker the bond
  - the shorter the bond, the fewer the electrons in it
  - the shorter the bond, the lower the bond dissociation energy

Study Guide for Content Mastery

Chemistry: Matter and Change • Chapter 9



Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

### Section 12 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write *true* or *false*.

- false** 1. Binary molecular compounds are generally composed of a metal and a nonmetal.
- false** 2. The second element in the formula of a binary compound is named using the suffix *-ite*.
- false** 3. The prefix *mono-* indicates three atoms.
- true** 4. The prefix *hepta-* indicates six atoms.
- true** 5. In naming the first element in a formula, the prefix *mono-* is not used.
- true** 6. For binary acids, the hydrogen part of the compound is named using the prefix *hydro-*.
- false** 7. An crystal contains only two elements.
- false** 8. If the name of the cation of an ionic salt is not the cation name, the salt name contains the suffix *-ate*.

In your textbook, read about naming molecular compounds and oxoacids.

For each item in Column A, write the letter of the matching item in Column B.

## Column A

5. \_\_\_\_\_ 6. CO
7. \_\_\_\_\_ 10. CO<sub>2</sub>
8. \_\_\_\_\_ 11. H<sub>2</sub>CO<sub>3</sub>
9. \_\_\_\_\_ 12. NO<sub>2</sub>
10. \_\_\_\_\_ 13. N<sub>2</sub>O<sub>4</sub>
11. \_\_\_\_\_ 14. HNO<sub>3</sub>
12. \_\_\_\_\_ 15. HNO<sub>2</sub>
13. \_\_\_\_\_ 16. HBr
14. \_\_\_\_\_ 17. HBrO<sub>3</sub>

## Column B

- a. hydrobromic acid
- b. nitrogen trioxide
- c. carbon monoxide
- d. nitrous acid
- e. carbonic acid
- f. nitric acid
- g. carbonic acid
- h. bromic acid
- i. carbon dioxide

50 Chemistry: Matter and Change • Chapter 9

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## **Table of Contents Teaching Transparency Master Answer Key**

1. Understanding the eBook Teaching Transparency Master Answer Key
  - The Rise of Digital Reading Teaching Transparency Master Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Teaching Transparency Master Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Teaching Transparency Master Answer Key
  - User-Friendly Interface

4. Exploring eBook Recommendations from Teaching Transparency Master Answer Key
  - Personalized Recommendations
  - Teaching Transparency Master Answer Key User Reviews and Ratings
  - Teaching Transparency Master Answer Key and Bestseller Lists
5. Accessing Teaching Transparency Master Answer Key Free and Paid eBooks
  - Teaching Transparency Master Answer Key Public Domain eBooks
  - Teaching Transparency Master Answer Key eBook Subscription Services
  - Teaching Transparency Master Answer Key Budget-Friendly Options
6. Navigating Teaching Transparency Master Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Teaching Transparency Master Answer Key Compatibility with Devices
  - Teaching Transparency Master Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Teaching Transparency Master Answer Key
  - Highlighting and Note-Taking Teaching Transparency Master Answer Key
  - Interactive Elements Teaching Transparency Master Answer Key
8. Staying Engaged with Teaching Transparency Master Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Teaching Transparency Master Answer Key
9. Balancing eBooks and Physical Books Teaching Transparency Master Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Teaching Transparency Master Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Teaching Transparency Master Answer Key
  - Setting Reading Goals Teaching Transparency Master Answer Key
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Teaching Transparency Master Answer Key
  - Fact-Checking eBook Content of Teaching Transparency Master Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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