

EDTA Titration for Determination of Calcium and Magnesium

Before attempting this experiment, you may need to consult the section in the textbook dealing with complexometric titrations. In this procedure a stock solution of EDTA will be standardized against primary standard CaCO_3 . This standardized EDTA solution is then used to determine water "hardness" – the total calcium and magnesium content, typically expressed as equivalent calcium carbonate mass concentration. While both metals are nutrients that are needed for good human health, high concentrations of these cations (Ca^{2+} and Mg^{2+}) in water can cause deposits to form in bathrooms and kitchens, and cause the formation of "soap scum" – clogging pipes and decreasing the effectiveness of the cleanser. Therefore, very hard water is undesirable for use as tap water. Guidelines for classification of water as provided by the U.S. Geologic Survey are listed in the table below. The USGS website has more information about water hardness and alkalinity if you are interested. We're fortunate in that most of the water in the Portland area is soft.

Table 1 U.S. Geological Survey Classification of Water Hardness

(Source: <http://water.usgs.gov/wmg/hardness-alkalinity.html>)

CaCO_3 (mg/L)	Hardness
0-60	soft
61-120	moderately hard
121-180	hard
>180	very hard

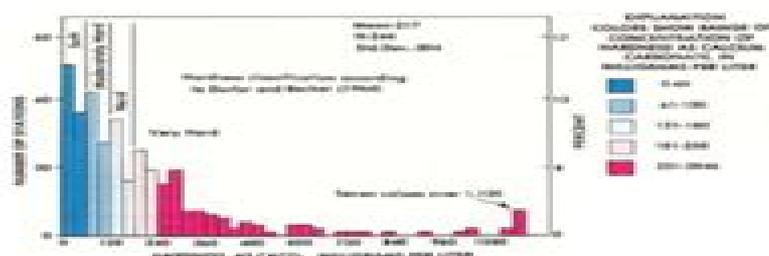


Figure 1 Hardness classification from Briggs et al. 1977

Briggs, J.C., and Ficke, J.F., 1977, Quality of Rivers of the United States, 1975 Water Year -- Based on the National Stream Quality Accounting Network (NASQAN): U.S. Geological Survey Open-File Report 78-200, 436 p.

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Food Analysis Laboratory Manual S. Suzanne Nielsen, 2010-03-20 This second edition laboratory manual was written to accompany Food Analysis Fourth Edition ISBN 978 1 4419 1477 4 by the same author The 21 laboratory exercises in the manual cover 20 of the 32 chapters in the textbook Many of the laboratory exercises have multiple sections to cover several methods of analysis for a particular food component of characteristic Most of the laboratory exercises include the following introduction reading assignment objective principle of method chemicals reagents precautions and waste disposal supplies equipment procedure data and calculations questions and references This laboratory manual is ideal for the laboratory portion of undergraduate courses in food analysis

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A Brief Introduction to Quantitative Chemical Analysis Robert Blanchard Fischer, Dennis G. Peters, 1969

Selected Technical Publications United States. Food and Drug Administration, 1966

Inorganic Titrimetric Analysis Walter Wagner, Clarence Joseph Hull, 1971

Treatise on Titrimetry: Inorganic titrimetric analysis, by W. Wagner and C. J. Hull Joseph Jordan, 1971

Report - US Army Medical Research Laboratory Army Medical Research Laboratory, 1947

Chemical Analysis Herbert August Laitinen, 1960 Equilibrium and activity Acid base equilibria in water Acid base equilibria in nonaqueous solvents Applications of acid base titrations Solubility of precipitates the formation of precipitates Colloidal properties of precipitates Aging of precipitates Contamination of precipitates Thermal decomposition and volatilization

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