

PHYSICAL SCIENCES - GRADE 10

1 Revision of matter (Grade 9)

1.1 Matter and classification

- Matter is made up of particles (atoms or molecules) and it is the properties of the atoms or molecules that determine the characteristics and reactivity of that matter.
- The properties of matter include: strength; density; melting and boiling points; whether the material is brittle, malleable or ductile; whether it is magnetic or not; and its electrical and thermal conductivity.
- Elements and compounds are classified as pure because they contain particles that are all the same. Substances can be classified as pure – elements and compounds
 - or mixtures. Pure substances contain particles that are all the same. Elements and compounds can be represented by symbols and formulae. Some of the names, symbols or formulae of a number of common elements and compounds that you should have learnt are the following:
 - Elements: S for sulphur; C for carbon; P for phosphorus (note the spelling); H for hydrogen; O for oxygen; He for helium; N for nitrogen; F for fluorine; Mg for magnesium; Ca for calcium; Zn for zinc.
 - Compounds: H_2O for water; H_2SO_4 for sulphuric acid; HCl for hydrochloric acid; NaCl for sodium chloride (table salt); $CaCO_3$ for calcium carbonate; KNO_3 for potassium nitrate; Na_3PO_4 for sodium phosphate.
- You will notice that the naming of compounds follows these rules:
 - The metal or hydrogen is always named first, then the non-metal (potassium iodide: KI; hydrogen chloride: HCl (also called hydrochloric acid)).
 - The name of a non-metal ion of an element always ends in -ide (for instance, sodium chloride: NaCl).
 - The name of a non-metal complex ion (more than one element in the ion) usually ends in -ate or -ite (for instance, copper sulphate: $CuSO_4$; potassium sulphite: K_2SO_3). An exception to this is hydroxide (for instance, sodium hydroxide: NaOH).
- Mixtures consist of different substances, and can be homogeneous (have the same 'look' throughout) or heterogeneous (you can see the different substances). They can be separated in various ways, for instance filtering, sieving, chromatography, distilling or dissolving one constituent.
- Some of the properties of metals are that they conduct electricity and heat, and they are malleable and ductile.
- Non-metals are non-conductors of electricity and heat. They are gases, liquids or brittle solids, and bond with each other.
- Metalloids are elements on the border between metals and non-metals, and they have both metal and non-metal properties.

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