

Stoichiometry: Mass – Mole Worksheet

Solve the following stoichiometry problems and type in the correct answer with units.
 (for example: 3.45 mol or 4.62 g) Round the masses from the periodic table to the nearest whole number. Round all answers to 2 decimal places.

1. Based on the following chemical equation, please answer the following questions



- a. Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₂.

Answer: _____ $\frac{\text{g FeCl}_2}{\text{g FeCl}_2} \left| \frac{1 \text{ mol FeCl}_2}{\text{g FeCl}_2} \right| \frac{\text{mol NaCl}}{\text{mol FeCl}_2} =$

- b. Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

Answer: _____ $\frac{\text{g Na}_2\text{SO}_4}{\text{g Na}_2\text{SO}_4} \left| \frac{1 \text{ mol Na}_2\text{SO}_4}{\text{g Na}_2\text{SO}_4} \right| \frac{\text{mol Fe}_2(\text{SO}_4)_3}{\text{mol Na}_2\text{SO}_4} \left| \frac{\text{g Fe}_2(\text{SO}_4)_3}{1 \text{ mol Fe}_2(\text{SO}_4)_3} \right| =$

- c. Calculate the amount in grams of FeCl₂ produced with 95g of Fe(SO₄)₂.

Answer: _____ $\frac{\text{g Fe (SO}_4)_2}{\text{g Fe (SO}_4)_2} \left| \frac{1 \text{ mol Fe (SO}_4)_2}{\text{g Fe (SO}_4)_2} \right| \frac{\text{mol FeCl}_2}{\text{mol Fe (SO}_4)_2} \left| \frac{\text{g FeCl}_2}{1 \text{ mol FeCl}_2} \right| =$

- d. Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl.

Answer: _____ $\frac{\text{g NaCl}}{\text{g NaCl}} \left| \frac{1 \text{ mol NaCl}}{\text{g NaCl}} \right| \frac{\text{mol Fe}_2(\text{SO}_4)_3}{\text{mol NaCl}} =$

2. How many moles of CaCO₃ can be produced if we make 12.3 g of Ca(HCO₃)₂ reacts with CaCO₃?



Answer: _____ $\frac{\text{g Ca (HCO}_3)_2}{\text{g Ca (HCO}_3)_2} \left| \frac{1 \text{ mol Ca (HCO}_3)_2}{\text{g Ca (HCO}_3)_2} \right| \frac{\text{mol CaCO}_3}{\text{mol Ca (HCO}_3)_2} =$

Moles And Stoichiometry Packet Answers

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Table of Contents Moles And Stoichiometry Packet Answers

1. Understanding the eBook Moles And Stoichiometry Packet Answers
 - The Rise of Digital Reading Moles And Stoichiometry Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Moles And Stoichiometry Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Moles And Stoichiometry Packet Answers
 - Personalized Recommendations
 - eBook Platform User Reviews and Ratings
 - eBook Platform Bestseller Lists

5. Accessing Moles And Stoichiometry Packet Answers Free and Paid eBooks
 - Moles And Stoichiometry Packet Answers Public Domain eBooks
 - Moles And Stoichiometry Packet Answers eBook Subscription Services
 - Moles And Stoichiometry Packet Answers Budget-Friendly Options
6. Navigating Moles And Stoichiometry Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Moles And Stoichiometry Packet Answers Compatibility with Devices
 - Moles And Stoichiometry Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Moles And Stoichiometry Packet Answers
 - Highlighting and Note-Taking Moles And Stoichiometry Packet Answers
 - Interactive Elements Moles And Stoichiometry Packet Answers
8. Staying Engaged with Moles And Stoichiometry Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Moles And Stoichiometry Packet Answers
9. Balancing eBooks and Physical Books Moles And Stoichiometry Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Moles And Stoichiometry Packet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Moles And Stoichiometry Packet Answers
 - Setting Reading Goals Moles And Stoichiometry Packet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Moles And Stoichiometry Packet Answers
 - Fact-Checking eBook Content of Moles And Stoichiometry Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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