

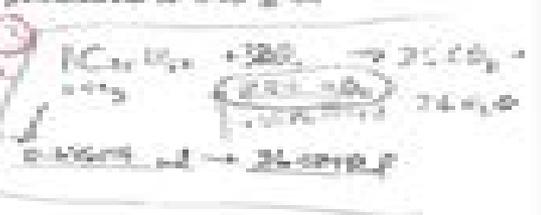
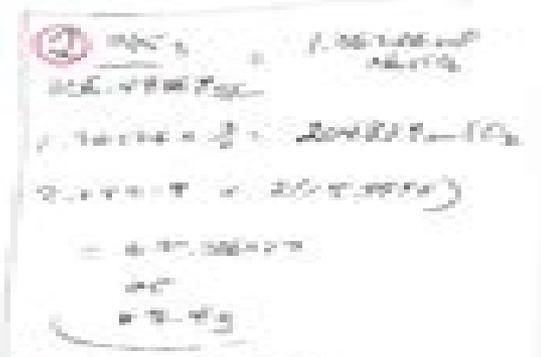
KEY

1. Complete and classify the following partial equations

- a) $1 \text{ Fe} + 3 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ FeSO}_4 + 3 \text{ H}_2$ (single repl)
- b) $\text{NaI} + \text{Cl}_2 \rightarrow$
- c) $1 \text{ Zn} + 1 \text{ Pb(NO}_3)_2 \rightarrow 1 \text{ Zn(NO}_3)_2 + 1 \text{ Pb}$ (single repl)
- d) $\text{Al(OH)}_3 + \text{HCl} \rightarrow$
- e) $2 \text{ Cu} + \text{O}_2 \rightarrow 2 \text{ CuO}$ $4 \text{ Cu} + 1 \text{ O}_2 \rightarrow 2 \text{ Cu}_2\text{O}$ (synthesis)
- f) $\text{Al}_2\text{O}_3 \rightarrow$
- g) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ (synth)
- h) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- i) $1 \text{ Mg} + 2 \text{ HNO}_3 \rightarrow 1 \text{ Mg(NO}_3)_2 + 1 \text{ H}_2$ (s.r.)
- j) $\text{Na} + \text{H}_2\text{O} \rightarrow$
- k) $3 \text{ NaOH} + \text{H}_3\text{PO}_4 \rightarrow 1 \text{ Na}_3\text{PO}_4 + 3 \text{ H}_2\text{O}$ (acid base)
- l) $\text{H}_2\text{SO}_4 + \text{Al(OH)}_3 \rightarrow$
- m) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 2 \text{ CO}_2 + 3 \text{ H}_2\text{O}$ (combustion)
- n) $\text{AlCl}_3 \rightarrow$
- o) $2 \text{ KOH} + \text{H}_2\text{SO}_4 \rightarrow 1 \text{ K}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$ (acid base)
- p) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- q) $1 \text{ NaOH} + \text{HCl} \rightarrow 1 \text{ NaCl} + 1 \text{ H}_2\text{O}$ (acid base)
- r) $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
- s) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ (combustion)
- t) $\text{H}_2\text{O} \rightarrow$
- u) $2 \text{ HNO}_3 + \text{CuCO}_3 \rightarrow 1 \text{ Cu(NO}_3)_2 + 1 \text{ CO}_2$ (s.r.)
- v) $\text{CH}_4 + \text{O}_2 \rightarrow$
- w) $1 \text{ Pb} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ PbSO}_4 + 1 \text{ H}_2$ (s.r.)
- x) $\text{KI} + \text{Br}_2 \rightarrow$
- y) $1 \text{ Ba(OH)}_2 + 2 \text{ HNO}_3 \rightarrow 1 \text{ Ba(NO}_3)_2 + 2 \text{ H}_2\text{O}$ (acid base)
- z) $\text{Pb} + \text{Cu(NO}_3)_2 \rightarrow$



1
2
3
4
5



- 2. Given: $1 \text{ NaClO}_3 \rightarrow 1 \text{ NaCl} + 3 \text{ O}_2$ calculate the mass of oxygen gas produced if 143 g of NaClO_3 is decomposed. **65.4 g**
- 3. Calculate the mass of oxygen gas required to burn 242 g of wax ($\text{C}_{21}\text{H}_{44}$). **254 g**
- 4. If 1.40 moles of butane (C_4H_{10}) is burned
 - a) calculate the number of moles of oxygen gas required. **9.8 moles O_2**
 - b) calculate the number of moles of carbon dioxide produced. **5.6 moles CO_2**
 - c) calculate the mass of water produced. **10.8 moles H_2O**
- 5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.
 - a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. **27.4 g Al**
 - b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. **192 g Zn**
- 6. Given: $\text{Cu} + 2 \text{ AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{ Ag}$
 - a) How many grams of Cu are required to react with 3.00 g of silver nitrate? **0.544 g Cu**
 - b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? **1.70 g Ag**

Nelson Chemistry 30 Answer Key

Ensheng Dong



Nelson Chemistry 30 Answer Key:

Nelson Chemistry 30 Answer Key Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Nelson Chemistry 30 Answer Key**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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