

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the $[H_2]$, which will increase the rate, but has no effect on k
 - Due to Arrhenius equation, changing temperature changes the value of k .
 - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/L · s
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- What is the order with respect to: P_4 0
 H_2 1
- Write the rate law for this reaction. $rate = k[H_2]$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 s^{-1}$**

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/L · s
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to: SO_2 2
 O_3 0
- Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 molL^{-2} s^{-1}$**

8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. **R**
- What is the order with respect to each reactant? **A_2 1st; B_2 1st**

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M Tight



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Table of Contents Practice Problems Solutions Kinetics And Equilibrium

1. Understanding the eBook Practice Problems Solutions Kinetics And Equilibrium
 - The Rise of Digital Reading Practice Problems Solutions Kinetics And Equilibrium
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Problems Solutions Kinetics And Equilibrium
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Problems Solutions Kinetics And Equilibrium
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Problems Solutions Kinetics And Equilibrium
 - Personalized Recommendations
 - Practice Problems Solutions Kinetics And Equilibrium User Reviews and Ratings
 - Practice Problems Solutions Kinetics And Equilibrium and Bestseller Lists
5. Accessing Practice Problems Solutions Kinetics And Equilibrium Free and Paid eBooks
 - Practice Problems Solutions Kinetics And Equilibrium Public Domain eBooks

- Practice Problems Solutions Kinetics And Equilibrium eBook Subscription Services
- Practice Problems Solutions Kinetics And Equilibrium Budget-Friendly Options
- 6. Navigating Practice Problems Solutions Kinetics And Equilibrium eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Problems Solutions Kinetics And Equilibrium Compatibility with Devices
 - Practice Problems Solutions Kinetics And Equilibrium Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Problems Solutions Kinetics And Equilibrium
 - Highlighting and Note-Taking Practice Problems Solutions Kinetics And Equilibrium
 - Interactive Elements Practice Problems Solutions Kinetics And Equilibrium
- 8. Staying Engaged with Practice Problems Solutions Kinetics And Equilibrium
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Problems Solutions Kinetics And Equilibrium
- 9. Balancing eBooks and Physical Books Practice Problems Solutions Kinetics And Equilibrium
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Problems Solutions Kinetics And Equilibrium
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Problems Solutions Kinetics And Equilibrium
 - Setting Reading Goals Practice Problems Solutions Kinetics And Equilibrium
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Problems Solutions Kinetics And Equilibrium
 - Fact-Checking eBook Content of Practice Problems Solutions Kinetics And Equilibrium
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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